CHEMISTRY

BOOKS - UNITED BOOK HOUSE

MODEL QUESTION PAPER 16

Exercise

1. No of Subshell under principle quantum no 'n' :

A. n^2

B. n-1

C. 2n

D. n

Answer:

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2. Covalency is maximum when a bond is formed between—

A. Atoms of same clement

B. Size of the atoms are almost equal

C. Maximum differences of elctro positivity

D. Having same electronic Configuration.

Answer:

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3. Reversible process is that process in Which

A. Surroundings transformed into system

B. System Spontaneously converts to surronundings

C. There is no boundary between system and surroundings

D. Always there.exist equilibrium between system and surroundings.

Answer:

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4. The value of compressibility factor of an ideal gas:

A. 1

B. -1

C. 2

D. 0.7

Answer:

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5. For the reaction $3A \rightarrow 2B$ the rate $D\frac{B}{d}A$ is equal to

A. $-\frac{2}{3}\frac{dA}{dt}$

B. $\frac{3}{2}\frac{dA}{dt}$

C. $-\frac{1}{3}\frac{dA}{dt}$

D. $3\frac{dA}{dt}$

Answer:

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6. Which pair shows diagonal property

A. Al& Be

B. Li &'Na

C. Na & Mg

D. Al& Mg.

Answer:

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7. Which among the following having highest ionic Conductivity

A. $Li^{\oplus}$

B. $Na^{\oplus}$

C. $K^{\oplus}$

D. $Cs^{\oplus}$

Answer:

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8. Which pair shows geometrical isomerism—

A. d & l lactic acid

B. Maleic & fumeric acid

C. Active & meso tartasic acid

D. Acetone & 2-propanol.

Answer:

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9. Which functional group has least priority-

A. - OH

B. -CHO

C. -O-

D. -COOH.

Answer:

10. Correct order of halogenation reaction of hydrocarbon—

A. Chlorination > Bromination > Idoination

B. Bromination > Chlorination > Iodination

C. Chlorination > Iodination > Bromination

D. Bromination > Iodination > Chlorination.

Answer:

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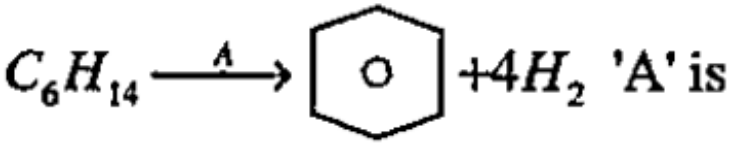
11. Find A in the figure:



'A' is

A. $LiAIH_4$

B. $K_2Cr_2O_7ConcH_2SO_4$

C. Na-ether

D. Cr_2O_3 / Al_2O_3

Answer:

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12. Which one is non biodegradable

A. Green vegetables

B. Gamaxine

C. Cow-dunk

D. Dead-body.

Answer:

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13. How many significant figure in 0.013200?

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14. What is lanthanide?

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15. What is stair step diagonal?

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16. WHat is thermodynamic equilibrium?

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17. What is octane number?

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18. Give an example of antiknock compound

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19. Calculate the number of H & O atoms in 90 gm water$at 4^{\circ} C$.

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20. State with example law of Reciprocal proportion.

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21. Calculate the wave length of photon having energy 2'ev.

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22. Find the first ionization energy of hydorgen atom given $E_1 = -13.38ev$

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23. Why the 1st Ionisation potential of N_2 is higher than O_2?

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24. Electron affinity of inert elements are positive. Why?

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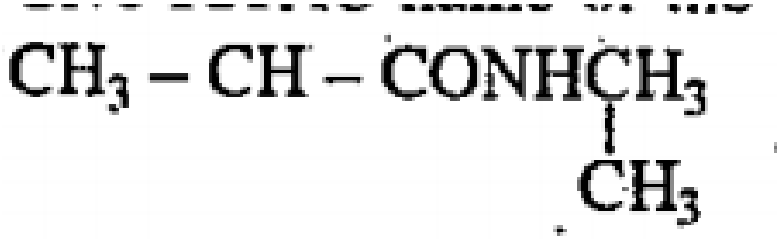
25. Give-IUPAC name of the following organic Compound



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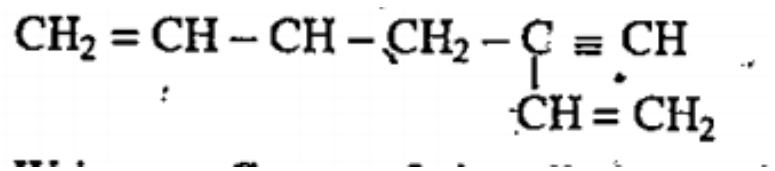
26. Give-IUPAC name of the following organic Compound



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27. Write two Causes of air pollution on State one. remidy of it.

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28. Write de Broglie equation for wave,particle duality.

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29. Show that P orbital can accomodate at least 6 electrons.

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30. What do you mean by lanthanoid Contraction?

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31. Arrange mg, A1, Si and Na in the increasing order of their lonisation potentials.

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32. What is electronegetivity? State with reason its change accross a period.

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33. Which one shows higher Covalent Character LiCl or Lil? and why?

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34. What is hybridisation? Explain with example sp^2 hybridisation.

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35. Write Graham's law of diffusion. Write the S.I'unit of viscosity?

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36. What is Capillery action? Give molecular interpretation of surface tension

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37. Calculate the final pressure when 2 molesof He at N..T.P. is compressed to 10 lit. Given $\gamma = 1.61$

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38. Bond dissociation energy ofC-C, C = C, H-H&C-Hare 300, 550,350, 400 Kj mol^{-1} Calculate heat of hydorgenation of ethylene.

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39. Find the oxidation number of * marked element in the following Compound:- $\left[Cr(NH_3)_6\right]Cl_3$

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40. Calculate the oxidation no of * marked atom:*Fe (CO)_5`

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41. Balance the euqation by oxidation no method:-

$$SnCl_2 + HCl + O_3 \rightarrow SnCI_4 + H_2O$$

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42. What is 'syn gas'? Why it is called syn gas?

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43. Can H_2O_2 Solution be concentrated by heating?

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44. What is volume strength? Which one is more powerful 10 volume and 10% H_2O_2

solution.

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45. Write the drawbacks of solveyprocess for synthesis of Na_2C0_3

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46. Write the IUPAC name of :-

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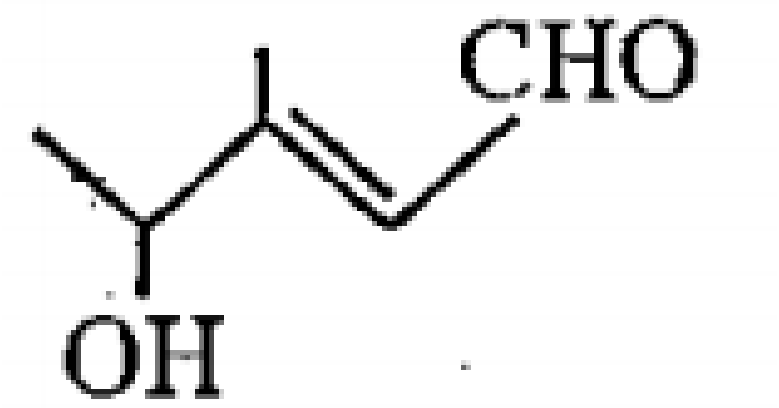
47. Write the IUPAC name of :-



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48. What is stereoganic Centere?

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49. Express the equilibrium constant of the reaction in the forms of $K_p$and K_c- for the reaction, $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$.

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50. State two differences between physical & Chemical equlibrium state the effect of temperature for the equilibrium. $ICe \rightarrow water$

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51. Discuss the effect of heat on the equilibrium

$$N_2(g) + O_2(g) = 2NO(g) - 44KCal.$$

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52. Why Boron has hight melting point? How would you -synthesis diborane? What is Zir Con?

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53. Explain why $\mathbb{C}l_4$ is not hydrolysed while $SiCl_4$ is hydroglysed

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54. What happens when:- $HCOOH$ is treated with Conc. H_2SO_4

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55. What is flint glass?

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56. How would you Carry out the following Conversion:Isopropyl bromide $\rightarrow$ 2, 3-dimethyl butane

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57. Do the following conversions:

Acetelene from Methane

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58. What is Marsh gas.

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