



CHEMISTRY

BOOKS - UNITED BOOK HOUSE

MODEL QUESTION PAPER 19

Exercise

1. Which one among the following disobeys Bohr theory

A. H

B. He^{\oplus}

C. Li_2^+

D. H^{\oplus}

Answer:



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2. Type of bond expected between two atoms having electro-negativities 1.0 & 3.8

A. Mejalic

B. Covalent

C. Co-ordinate Covalency

D. electrovalent.

Answer:



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3. Structure of NH_3 —

A. Tetrahedral

B. Pyramidal

C. Linear

D. Triangular

Answer:



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4. Value of R in $Cal / k. Mole-$

A. 8.31

B. 8.31×10^7

C. 0.082

D. 1.987. 5

Answer:



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5. Relation between ΔH & ΔU

A. $\Delta H = \Delta U + \Delta nRT$

B. $\Delta H = \Delta U - P\Delta V$

C. $\Delta H = \Delta U - P\Delta VS$

$$D. \Delta U = \Delta H + P\Delta V$$

Answer:



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6. Change of enthalpy in exothermic reaction

A. Positive

B. Negative

C. Zero

D. infinity.

Answer:



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7. Velocity Constant for a reaction at 290.K is given by $3.2 \times 10^{-3} S^{-1}$. What will be its value at 300 K—

A. 6.4×10^{-3}

B. 3.2×10^{-4}

C. 9.6×10^{-3}

D. 1.28×10^{-2}

Answer:



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8. Which one of the following cation has highest polarising Power

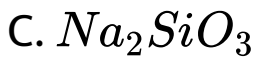
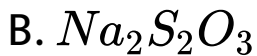
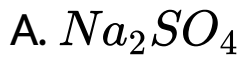


Answer:



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9. Water glass is

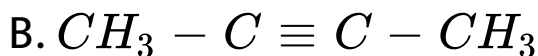
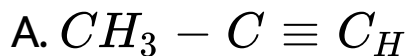


Answer:

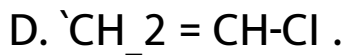
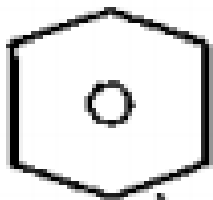


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10. Which one of the following shows acidic nature—



C.



Answer:



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11. Largest C -C bond preetn in—

- A. propane
- B. ethene
- C. Acetylene
- D. Benzene.

Answer:



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12. The reaction between $AgNO_3$ and C_2H_2 is

A. oxidation

B. Reduction

C. Acidbase

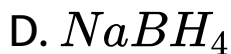
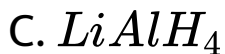
D. Substitution.

Answer:



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13. Which one is found in Bayer's reagent—



Answer:



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14. The lung disease caused by silica is—

A. Minamata

B. Itai - Itai

C. Silicosis

D. Black foot

Answer:



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15. Calculate the mass of 2.24 liter CO_2 at N.T.P



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16. No of Ca^{2+} ion present in 1 gm-ion Ca^{2+} is?



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17. Give an example of actionid.



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18. What is state function?



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19. What type of function entropy is?



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20. Give an example of Metamerism



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21. State laws of reciprocal proportion.



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22. Why law of reciprocal proportion is called equivalent ratio rule?



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23. Explain with proper-example pauli exclusion principle.



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24. State two postulates of Bohr atom model



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25. State with equation effect of temperature on Boric Acid.



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26. Why BF_6^{3-} does not exist, explain.



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27. Which one is weak acid and why?

ICH_2COOH , $BrCH_2COOH$, $ClCH_2COOH$

.



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28. Which is most basic and why? CH_3NH_2 ,
 $(CH_3)_2NH$, $(CH_3)_3N$



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29. What is particulate matter? Give two example'



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30. Write two defects of Bohr atom model.
Concept of which quantum number comes from Sommerfeld theory?



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31. Derive the equation for radius of n th shell from Bohr theory.



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32. What is d-block element? Give two example



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33. Why the S-block element does not respond to flame test? State two properties of element which are not peiodic.



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34. State reason:- Bond angle of PF_3 is greater than NH_3



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35. State reason:- CO_2 is linear but not SO_2



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36. Write two differences between σ & π bond.

What is antibonding molecular orbital?



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37. State Daltons partial pressure laws and derive partial pressure and total pressure relation.



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38. Calculate the partial pressure of SO_2 in a mixture with 60% CO_2 and total pressure being 4 atm.





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39. What,do you mean by enthalpy of a system? Write two characteristic of it.



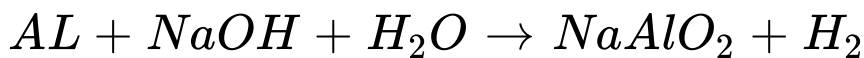
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40. Calculate, ΔU , W and Q when an ideal gas is compressed at 5 atm pressure from 25 litre to 5 litre at a constant temp of $27^\circ C$.



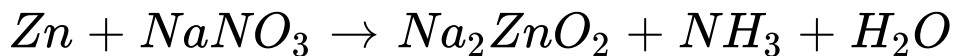
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41. Balance the equation by oxydation no. method :-



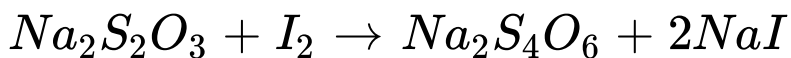
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42. Balance by ion electron method :



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43. Balance the equation by oxydation no. method :-



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44. Write with balanced equation what happens when:- H_2O_2 is passed through acidified $KMnO_4$ soln



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45. Write with balanced equation what happens when: H_2O_2 is passed through acidified $K_2Cr_2O_7$ soln



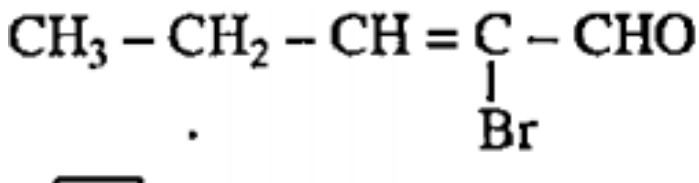
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46. Which one is more soluble in water and why $MgCO_3$ & $CaCO_3$. What is hydrolysis?



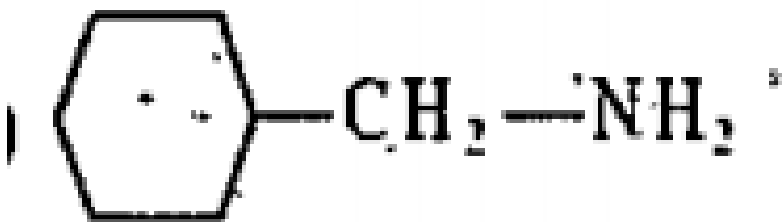
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47. Write the IUPAC name of the following compound:-



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48. Write the IUPAC name of the following compound:-





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49. Write the IUPAC name of the following compound:- $CH_2 = C = CH_2$



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50. $aA + bB = cC + dD$ If the equilibrium Constant for the reaction be K . get the equilibrium constant for the reaction $XaA + XbB + XcC + XdD$.



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51. Write La Chatterlicr principle. Applying this Discuss the effect of heat on the equilibrium of an exothermic reaction.



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52. How would you prepare:- Boric Acid from Borax.



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53. How would you prepare:- Borax from Colemanite



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54. State one use of Boric acid



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55. What happens when:-CO is passed through ammoniacal Silver nitrate solution



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56. what happens when CO is passed through Cuprous Chloride solution.



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57. State one use of CO



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58. Write short note on:- Ozonolysis reaction.

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59. Write short note on:-Wurtz reaction

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60. How would you convert:-Acetylene \rightarrow 2 -
Chloro propene

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61. How would you convert: Ethylene \rightarrow
Glyoxal



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62. What is Lindler's Catalyst.



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