



CHEMISTRY

BOOKS - UNITED BOOK HOUSE

TIRTHAPATI INSTITUTION QUESTION PAPER

Exercise

1. Write one drawback of Dalton's atomic theory.



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2. Illustrate with example law of multiple proportion.



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3. In a hydrocarbon $y.c = 75$. vapour density of the compound .Find the empirical & molecular formula of the compound.



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4. State Bohr's theory.

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5. The uncertainty in the momentum of an electron is $10^{-5} \text{ kgms}^{-1}$. Calculate the uncertainty of its position.

$$[h = 6.62 \times 10^{-34} \text{ JS}]$$

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6. State Pauli's exclusion principle.

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7. State modern periodic law.



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8. Write IUPAC name of an elements having atomic no.116.



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9. Why the electron affinity of chlorine is higher than that fluorine?



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10. Write the general outer electronic configuration of f block elements.



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11. Explain the shape of $BeCl_2$ & NH_4^+ according to VSEPR theory.



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12. O_2 is paramagnetic, why?



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13. H_2O is liquid but H_2S is gas why?

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14. Define: Surface tension.

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15. Define: Critical temperature.

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16. Define: Avogadro's number.



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17. 0.64 g of gas at $0^{\circ}C$ and 75 cm Hg pressure occupy 225 ml volm. Find its molecular weight?



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18. Write the thermodynamic criteria for equilibrium condition and spontaneity.

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19. Heat of formation of CO_2 , H_2O & CH_4 are -100cal , -50cal & -70cal . Find heat of combustion of CH_4 .

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20. 2 mol of an ideal gas at 3 atm pressure is expanded from 10 lit to 20 lit. Find work done and change in internal energy.

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21. State Le-Chateliers principle.



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22. Find the pH of $0.1(N)HCl$ & $10^{-8}(N)HCl$ solution.



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23. Solubility of $Zn_3(PO_4)_2$ is 'S' mol / lit. Find its solubility product.



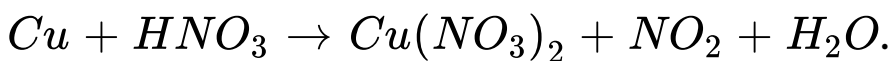
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24. Define oxidation in terms of electronic theory.

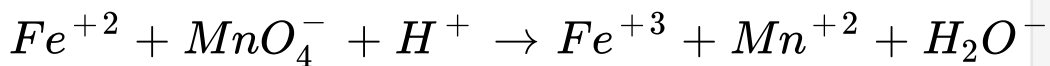
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25. Balance by oxidation number method



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26. Balance by ion-electron method.



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27. Name a metal which liberate H_2 gas .from acid and alkali.



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28. Write the principle of laboratory preparation of H_2O_2 with equation.



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29. Give one example of ionic hydride.



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30. Give one example of Co-valent Hydride.



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31. What is Buffer solution.



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