CHEMISTRY

AAKASH INSTITUTE ENGLISH

TEST 4

Example

1. The specific gravity of SO_3 vapour at0^oC and 76 cmHg in grams/ litre is

A. 5.6

B. 11.2

C. 3.57

D. 4.5

Answer:

2. The pH value of 0.20 M solution of CH_3NH_2 at 298 K is [Given that its ionisation constant K_b is 5.0×10^{-5}]

A. 12.23

B. 11.50

C. 10.90

D. 12.45

Answer:

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3. If the volume of a given mass of a gas at constant temperature becomes one third times, the pressure will be

A. 3p

B. p/3

C. 6p

D. 9p

Answer:

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4. A vessel at 1200 K contains $CO_2(g)$ at a pressure of 0.5 atm. Some of the CO_2 is converted into CO(g) on addition of graphite.

The value of equilibrium constant (K_p) for the process, if total pressure at equilibrium is 0.7 atm, is

A. 0.53 atm

B. 0.92 atm

C. 0.76 atm

D. 0.26 atm

Answer:

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5. Some solid NH_4HS is placed in flask containing 0.5 atm of NH_3. What would be the pressure of NH_3 and H_2S when equilibrium is reached.

$NH_4HS(g) \Leftrightarrow NH_3(g) + H_2S(g), K_p = 0.11$

A. $P_{NH_3} = 0.70atm, P_{H_2S} = 0.20atm$

B. $P_{NH_3} = 0.64atm, P_{H_2S} = 0.25atm$

C. $P_{NH_3} = 0.80atm, P_{H_2S} = 0.26atm$

D. $P_{NH_3} = 0.45atm, P_{H_2S} = 0.32atm$

Answer:

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6. If the volume of a given mass of a gas at constant temperature becomes three times, the pressure will be

A. 4p

B. 3p

C. p/3

D. 9p

Answer:

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7. Which of the following bonds has the highest bond energy?

A. $S = O$

B. $C \equiv N$

C. $C \equiv C$

D. $N \equiv N$

Answer:

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8. Dissociation constant (K_a) of formic acid and acetic acid are 2.5×10^{-4} and 0.5×10^{-5} respectively. The ratio of their relative strengths at the same concentration is

A. 2.09

B. 7.07

C. 4.26

D. 6.03

Answer:

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9. At t^oC , K_w for water is 6.4×10^{-13}. The pH of water at t^oC will be

A. 6.91

B. 5.90

C. 6.09

D. 5.15

Answer:

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10. An ideal gas has specific heat at constant volume C_v = (3/2)R. The gas is kept in closed vessel of volume 0.0080 m^3 at a temperature of 300 K and pressure 1.6×10^6 N/m^2. If 2.8×10^4 J of heat is supplied to the gas, then the final temperature of the gas will be

A. 981

B. 782

C. 737.54

D. 437.54

Answer:

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11. Degree of hydrolysis (α) for a salt of strong acid and weak base is :

A. $\sqrt{\frac{K_w}{K_b \times C}}$

B. $\sqrt{\frac{K_w \times K_b}{K_a}}$

C. $\sqrt{\frac{K_w \times C}{K_b}}$

D. $\sqrt{K_a \times K_b}$

Answer:

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12. An ideal gas has specific heat at constant volume C_v = (3/2)R. The gas is kept in closed vessel of volume 0.0090 m^3 at a temperature of 300 K and pressure 1.8×10^6 N/m^2. If

2.8×10^4 J of heat is supplied to the gas, then the final temperature of the gas will be

A. 675 K

B. 725 K

C. 690 K

D. 680 K

Answer:

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13. The favourable conditions for the reaction,

$2H(g) \Leftrightarrow H_2(g)$ are

A. Low temperature, low pressure

B. High temperature, low pressure

C. High temperature, high pressure

D. Low temperature, high pressure

Answer:

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14. At $27^\circ C$, K_p value for the reaction

$CaCO_3(s) \Leftrightarrow CaO(s) + CO_2(g)$, is 0.1 atm.

K_C value for this reaction is

A. $4x10^{-3}$

B. $6x10^{-3}$

C. $2x10^{-3}$

D. $9x10^{-3}$

Answer:

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15. Which of the following is not a conjugate acid-base pair?

A. NH_4^+, NH_3

B. H_3O^+, OH^-

C. F^-, HF

D. $CH_3 - COOH, CH_3 - COO^-$

Answer:

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16. One litre sample of hard water contains 0.006 mole $CaCl_2$. The minimum concentration of Na_2SO_4 needed for the precipitation of $CaSO_4$ is

(K_{sp} for $CaSO_4$ is 3.0×10^{-6} at 25^oC)

A. 6.2×10^{-3}

B. 4.2×10^{-5}

C. 5×10^{-4}

D. 6×10^{-7}

Answer:

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17. In a buffer solution concentration of $(NH_4)_2SO_4$ and NH_4OH are 0.4 M and 0.8 M respectively, pH of the solution is $\left[K_a\left(NH_4^+\right) = 10^{-8}\right]$

A. 7.2

B. 8

C. 9.2

D. 10

Answer:

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18. At 25^oC, K_{sp} for AB_2 salt is equal to 8×10^{-6}. If the salt is 70% dissociated, then the solubility of AB_2 in mol/litre is

A. $S = \left(\frac{8 \times 10^{-6}}{0.98} \right)^{\frac{1}{3}}$

B. $S = \left(\frac{5 \times 10^{-6}}{0.76} \right)^{\frac{1}{3}}$

C. $S = \left(\frac{8 \times 10^{-6}}{0.98} \right)^{\frac{1}{2}}$

D. $S = \left(\frac{5 \times 10^{-6}}{0.76} \right)^{\frac{1}{2}}$

Answer:

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19. At constant temperature, vapour density of N_2O_4 is found to be 40. Percentage degree of dissociation of N_2O_4 at this constant temperature will be $N_2O_4(g) \Leftrightarrow 2NO_2(g)$

A. 0.17

B. $16.6\ \%$

C. $19.2\ \%$

D. 0.15

Answer:

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20. Which one of the following has the highest basic character?

A. Na_2O

B. K_2O

C. Rb_2O

D. Li_2O

Answer:

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21. Calculate the change in entropy $(\Delta_r S^o)$ at 298 K for the reaction in which urea is formed from NH_3 and $CO_2 2NH_3(g) + CO_2(g) \rightarrow$ NH_2CONH_2 + H_2O . The standard entropy of

$NH_2CONH_2(aq), CO_2(g), NH_3(g)$ and $H_2O(l)$ are 174.0, 213.7, 195 and 75 $JK^{-1}mol^{-1}$ respectively

A. $-354.7JK^{-1}mol^{-1}$

B. $-387JK^{-1}mol^{-1}$

C. $-390JK^{-1}mol^{-1}$

D. $-375JK^{-1}mol^{-1}$

Answer:

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22. At 25^oC, if the concentration of Ag^+ ion is 1.5×10^{-4} mol/L in the saturated solution of $Ag_2Cr_2O_7$, then solubility product of $Ag_2Cr_2O_7$ is

A. 6.25×10^{-16}

B. 5.1×10^{-16}

C. 1.69×10^{-12}

D. 7.2×10^{-14}

Answer:

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23. Among the following which salt do not undergo hydrolysis?

A. CH_3COONa

B. NaCN

C. Na_2CO_3

D. $NaNO_3$

Answer:

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24. If 2^oC rise in temperature takes place on the combustion of 1.0 g of methane gas in a bomb calorimeter (heat capacity = 500 $J/^oC$) at 27^oC, then the value of heat of combustion at constant pressure (in kJ) for one mole of methane gas is

A. -150

B. -165

C. -154

D. -169

Answer:

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25. In a system $A(s) \Leftrightarrow 2B(g) + 3C(g)$, if the concentration of C at equilibrium is increased by a factor of 2, it will cause the equilibrium concentrationof B to change to :

A. 2 times

B. $\frac{1}{2}$ times

C. $\frac{2\sqrt{2}}{3}$ times

D. $\frac{1}{2\sqrt{2}}$ times

Answer:

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26. 75 calories of heat is required to raise the temperature of 3 moles of an ideal gas at constant pressure from 30^oC to 35^oC. The amount of heat required in calories to raise the temperature of the gas through the range

$(50^oC$ to $55^oC)$ at constant volume is: (y= Cv/Cp =1.4)

A. 33cal

B. 65.26cal

C. 53.55cal

D. 75cal

Answer:

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27. At 100^oC and 1 atm pressure the density of water vapour is 0.0005970 g/cc.What is the molar volume and how does this compare with ideal gas volume ? What is the compressibilty factor ' Z '?

A. 1

B. 0.758

C. 0.985

D. 0.576

Answer:

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28. The standard Gibbs energy change for the conversion of oxygen to ozone at 300 K is [Given that K_p for the change $\left(\frac{3}{2}\right) O_2(g) \Leftrightarrow O_3(g)$ at 300 K is 2.5×10^{-30} and $\log_{10} 5 = 0.699$]

A. 170.38 kJ

B. 178.28 kJ

C. 182.48 kJ

D. 189.38 kJ

Answer:

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29. Prove that the pressure necessary to obtain $50\,\%$ dissociation of PCl_5 at $250^\circ C$ is numerically three times of K_p.

A. $K_p = \frac{9}{16}P$

B. $K_p = \frac{16}{9}P$

C. $K_p = \frac{2}{5}P$

D. $K_p = \frac{5}{2}P$

Answer:

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30. If enthalpy of neutralisation of HCl by NaOH is -58.84 kJ/mol and by NH_4OH is -52.26 kJ/mol, then enthalpy of ionisation of NH_4OH is

A. 6.58 kJ/mol

B. - 6.58 kJ/mol

C. 4.5 (kJ)/mol

D. - 4.5 kJ/mol`

Answer:

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31. A gas expands from 4.0 L to 4.5 L against a constant external pressure of 1 atm. The work done by the gas is (1 L-atm = 101.3 J)

A. $-62.25J$

B. $-50.65J$

C. $-52.25J$

D. $-60.25J$

Answer:

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32. Volume of 1.5 mole of a gas at 1 atm. pressure and 273K is

A. 22.4L

B. 33.6L

C. 11.2L

D. 10L

Answer:

33. In which of the following cases, the solution of sparingly soluble salt AB is unsaturated?

A. $\left[A^{+}\right]\left[B^{-}\right] < K_{sp}$

B. $\left[A^{+}\right]\left[B^{-}\right] > K_{sp}$

C. `[A^+][B^-]=K_(sp)

D. $\left[A^{+}\right]\left[B^{-}\right] \geq K_{sp}$

Answer:

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34. Consider the following equilibrium reactions at $25^{o}C$,

$A(g) + 3B(g) \Leftrightarrow 2C(g), K_{eq} = x$

$A(g) + D(g) \Leftrightarrow 2E(g), K_{eq} = y$

$B(g) + D(g) \Leftrightarrow F(g), K_{eq} = z$

then the value of equilibrium constant for the reaction $2C + 4D \Leftrightarrow 2E + 3F$, at 25^oC is

A. $\frac{yz^3}{x}$

B. $\frac{xy}{z^3}$

C. $\frac{xz}{y}$

D. $\frac{y^2z}{x}$

Answer:

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35. Among the following processes, select the process for which ΔS is positive.

A. $N_2(g) + O_2(g) \rightarrow NO(g)$

B. $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$

C. $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

D. $NaNO_2(s) + \frac{1}{2}O_2(g) \rightarrow NaNO_3(s)$

Answer:

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36. When He(g) is added in the following equilibrium at constant volume

$$N_2(g) + 3H_2(g) \Leftrightarrow 2NH_3(g)$$

Then select the correct statement.

A. Concentration of NH_3 increases

B. Concentration of NH_3 remains same

C. Concentration of N_2 decreases

D. Total number of moles decreases

Answer:

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37. The equilibrium constant K_{p1} and K_{p2} for the reactions $X \Leftrightarrow 2Y$ and $Z \Leftrightarrow P + Q$, respectively, are in the ratio of $1:9$. If the degrees of dissociation of X and Z are equal, then the ratio of total pressure at equilibria is

A. $1:6$

B. $1:1$

C. $1:3$

D. $1:9$

Answer:

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38. For a reversible adiabatic expansion of an ideal gas dp/p is equal to

A. $\gamma \frac{dv}{v}$

B. $\frac{dv}{v}$

C. $\left(\frac{\gamma}{\gamma - 1}\right) \frac{dv}{v}$

D. $-\gamma \frac{dv}{v}$

Answer:

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39. Among the following select the Lewis acid-base pair respectively

A. $AICl_3, BF_3$

B. BF_3, NH_3

C. CH_3NH_2, CH_3OH

D. $FeCI_3, ZnCI_2$

Answer:

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40. The degree of hydrolysis of 0.005 M aqueous solution of NaOCN will be (K_a for HOCN = 4.0×10^{-4})

A. 7.07×10^{-5}

B. 5.06×10^{-6}

C. 6.02×10^{-5}

D. 4.56×10^{-4}

Answer:

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41. The incorrect criterion for the spontaneity of a process is

A. $\Delta G_{system} < 0$

B. $\Delta S_{system} + \Delta S_{surrounding} > 0$

C. $\Delta S_{system} > \Delta S_{surrounding}$

D. $\Delta H_{system} < 0$ and $\Delta S_{system} > 0$

Answer:

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42. For the reaction at 300 K, $N_2O_4(g) \Leftrightarrow 2NO_2(g)$, the value of ΔG when concentration of each species is 5 mol/litre, will be

(Given: $(\Delta G^o_{fN_2O_4})$ = 120 kJ/mol and $(\Delta G^o_{fNO_2})$ = 70 kJ/mol)

A. 24 kJ

B. 30 kJ

C. 26 kJ

D. 29 kJ

Answer:

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43. 50 J heat is supplied to a gaseous system, simultaneously the gas expands against external pressure 2 atm by 0.8 L. The change in internal energy of the system will be

A. -120.8 J

B. -112.11 J

C. -125.4 J

D. -130.60 J

Answer:

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44. A basic buffer solution can be prepared by mixing the solution of

A. NaOH & HCl

B. HCN & HCl

C. NH_4CI & NH_4OH

D. NH_4CI & HCl

Answer:

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45. Among the following the strongest acid is

A. C_2H_5OH

B. HF

C. H_2O

D. CH_4

Answer:

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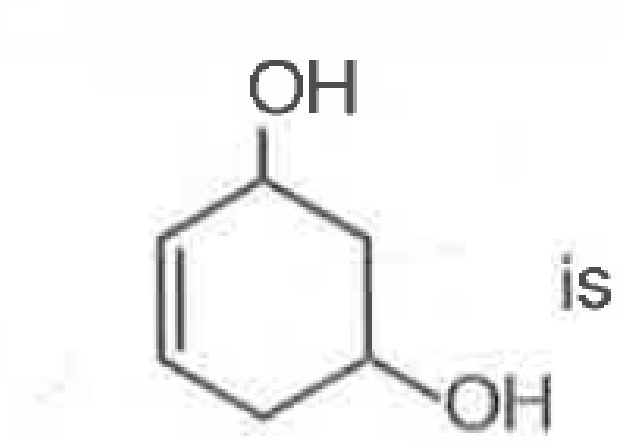
46. The correct IUPAC name of compound



is

A. Cyclohex-5-ene-1,3-diol

B. Cyclohex-4-ene-1,3-diol

C. Cyclohex-2-ene-1,5-dial

D. Cyclohex-3-ene-1,5-diol

Answer:

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47. For the reaction $C(s) + CO_2(g) \rightarrow 2CO(g), k_p = 63\text{atm}$ at 100 K. If at equilibrium $p_{CO} = 10p_{CO_2}$ then the total pressure of the gases at equilibrium is

A. 6.3atm

B. 6.93 atm

C. 0.63atm

D. 0.693atm

Answer:

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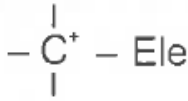
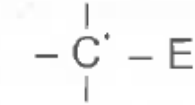
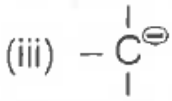
48. Which of the following is/are correctly matched?

(i)  – Electrophile, Lewis acid

(ii)  – Electron deficient

(iii)  – Nucleophile, Lewis base

(iv) >C: – Nucleophile

A. (i), (ii) & (iii) only

B. (ii), (iii) & (iv) only

C. (i) & (iii) only

D. (i), (ii), (iii) & (iv)

Answer:

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49. Total number of isomers of C_7H_8O containing benzene ring is

A. 3

B. 4

C. 5

D. 6

Answer:

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50. The total number of cyclic isomers possible for a hydrocarbon with the molecular formula C_4H_6 is .....

A. 4

B. 5

C. 6

D. 3

Answer:

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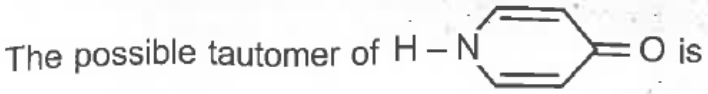
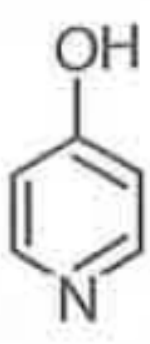
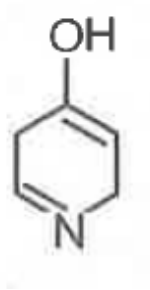
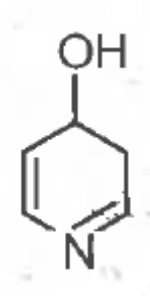
51. Complete the following reaction

The possible tautomer of  is

A. 

B. 

C. 

D. Tautomerism is not possible

Answer:

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52. Which of the following has highest —I effect?

A. —NO_2

B. —F

C. —OH

D. — NH_3^+

Answer:

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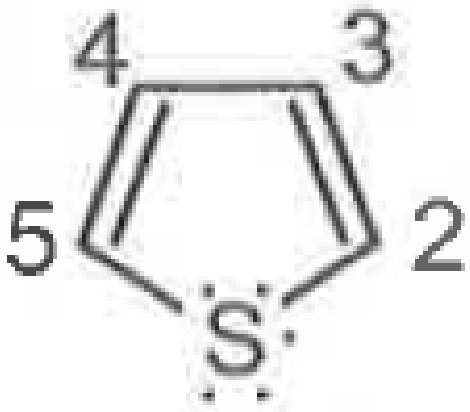
53. In thiophene, the positions at which electron density will be maximum are



A. 2, 3

B. 3, 4

C. 2, 5

D. 2, 4

Answer:

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54. Geometrical isomerism is not possible in

A. Allene

B. Aldoxime

C. Azobenzene

D. 1,2-Dichlorocyclohexane

Answer:

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55. The index of hydrogen deficiency in $C_{20}H_{20}O_2$ is

A. 11

B. 6

C. 9

D. 12

Answer:

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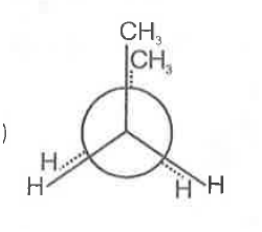
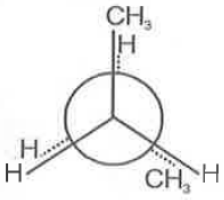
56. Which of the following is most stable conformer?

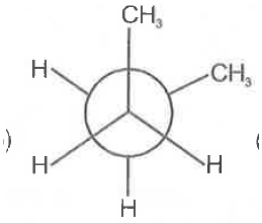
A.



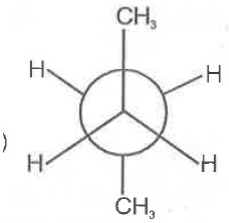
B.



C.



D.



Answer:

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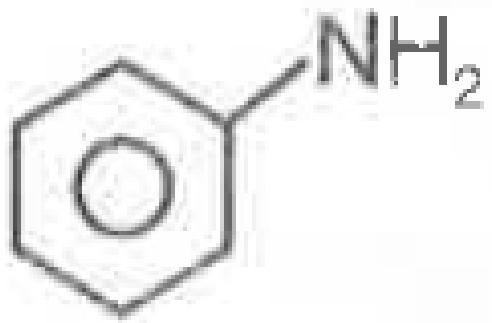
57. The correct order of basicity among the following is

(i) 

(ii)

(iii)

A. $(i) > (iii) > (ii)$

B. $(ii) > (i) > (iii)$

C. $(i) > (ii) > (iii)$

D. $(ii) > (iii) > (i)$

Answer:

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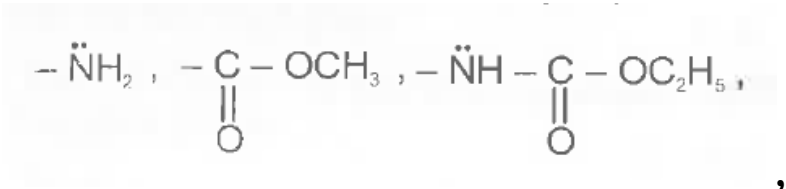
58. The mesomeric effects of the groups



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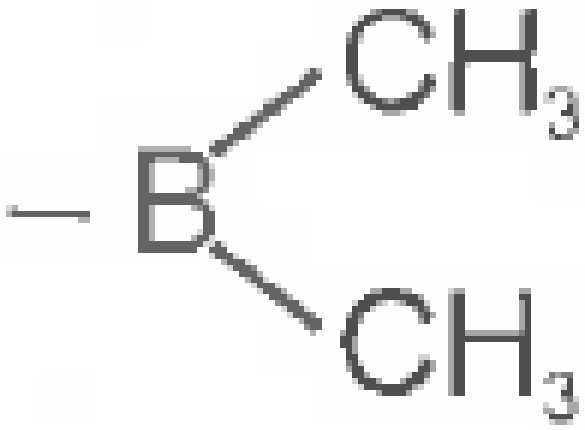

respectively is

A. + M, + M, — M, M

B. + M, — M, + M, -M

C. — M, — M, + M, M

D. + M, — M, — M, + M

Answer:

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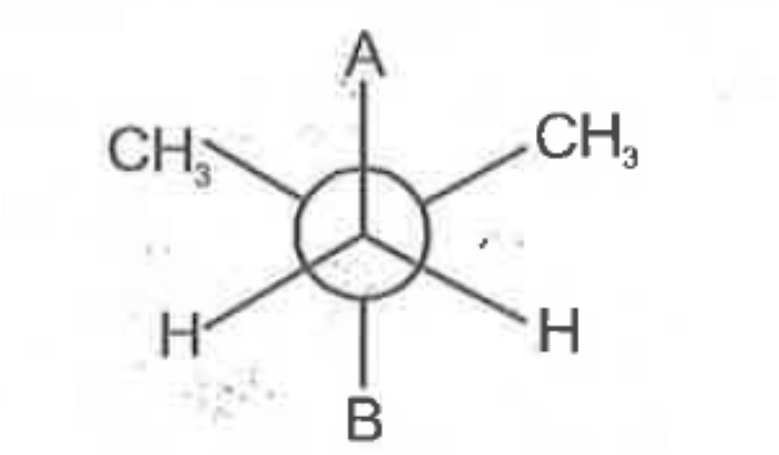
59. In the given Newmann's projection formula of 2,2-Dimethyl butane, A and B respectively can be



A. H and H

B. CH_3 and C_2H_5

C. CH_3 and H

D. CH_3 and CH_3

Answer:

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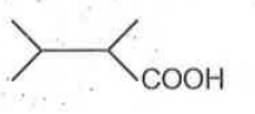
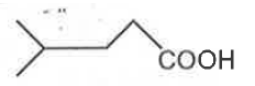
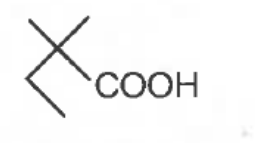
60. The acid which does not give isopentane on decarboxylation with soda lime, is

A. 

B. 

C. 

D. 

Answer:

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61. Calculate the percentage of free volume available in 1 mole gaseous water at 1 atm pressure and 373 K.

A. 101

B. 99.9

C. 90

D. 80

Answer:

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62. For the reaction $C(s) + CO_2(g) \Leftrightarrow 2CO(g)$, the partial pressure of CO_2 and CO both is 4.0 atm, respectively, at equilibrium. The K_p of the reaction is

A. 10atm

B. 4atm

C. 8atm

D. 0.4atm

Answer:

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63. For the reaction $C(s) + CO_2(g) \Leftrightarrow 2CO(g)$, the partial pressure of CO_2 and CO both is 5.0 atm, respectively, at equilibrium. The K_p of the reaction is

A. 20atm

B. 5atm

C. 15atm

D. 10atm

Answer:

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64. For the reaction $C(s) + CO_2(g) \Leftrightarrow 2CO(g)$, the partial pressure of CO_2 and CO is 4.0atm amd 8.0 atm, respectively, at equilibrium. The K_p of the reaction is

A. 16atm

B. 8atm

C. 32atm

D. 4atm

Answer:

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65. At $27^\circ C$, K_p value for the reaction

$CaCO_3(s) \Leftrightarrow CaO(s) + CO_2(g)$, is 2 atm.

K_C value for this reaction is

A. 8×10^{-2}

B. 4×10^{-2}

C. 6×10^{-2}

D. 4×10^{-3}

Answer:

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66. The term "Retardation factor" is related to the

A. Solvent extraction

B. Dumas method

C. Chromatography

D. Carius method

Answer:

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67. Which of the following sets of elements can be detected by Lassaigne's test?

A. N, H, P

B. N, Cl, P, S

C. N, Br, C, P

D. N, S, F, C

Answer:

68. In Carius method, 0.18 gram of an organic compound gave 0.11 gram of AgBr. The percentage of Br in the compound is

A. 34.0

B. 39.0

C. 26.0

D. 3.0

Answer:

69. The number of moles of sodium acetate required to produce 112 ml of ethane at STP by Kolbe's electrolysis is

A. 1×10^{-2}

B. 5×10^{-4}

C. 2×10^{-3}

D. 2×10^{-2}

Answer:

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70. Which of the following has highest boiling point?

A. 

B. 

C. 

D. 

Answer:

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71. Given the reaction at $967^{\circ}C$ and 1 atm.

$$CaCO_3(s) \Leftrightarrow CaO(s) + CO_2(g)$$

DeltaH=176 $kJmol^{-1}$, then DeltaE equals

A. 156.6KJ

B. 165.6KJ

C. 16.6KJ

D. 1.656KJ

Answer:

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72. The incorrect statement regarding ethylene is

A. It has higher heat of than other members of alkene series

B. Its degree of unsaturation is equal to one

C. Nodal plane of it bond of ethylene is a plane perpendicular to its molecular plane

D. C—C bond length is longer than C—H bond length

Answer:

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73. 90 calories of heat is required to raise the temperature of 3 moles of an ideal gas at constant pressure from 30^oC to 35^oC. The amount of heat required in calories to raise the temperature of the gas through the range (50^oC to 55^oC) at constant volume is: (γ= Cp/Cv =1.4)

A. 64.28cal

B. 74.68cal

C. 78cal

D. 90cal

Answer:

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74. Write expression for $K_p \text{and} K_c$ for the reaction $CaCO_3(S) \leftrightarrow CaO(S) + CO_2(g)$.

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75. Given the reaction at $937^{\circ} C$ and 1 atm.

$$CaCO_3(s) \Leftrightarrow CaO(s) + CO_2(g)$$

DeltaH=175 kJmol^{-1}, then DeltaE equals

A. 156.6KJ

B. 166.6KJ

C. 164.9KJ

D. 15.6KJ

Answer:

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76. A flask of volume 1 litre contains vapour of CH_3OH at a pressure of 1 atm and 25^oC. The flask was then evacuated till the final pressure dropped to 10^{-4} mm. Find the number of molecules of methyl alcohol left in the flask.

A. 2.24×10^{15}

B. 2.24×10^{12}

C. 4.24×10^{15}

D. 3.24×10^{15}

Answer:

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77. Write the name of the gas evolved when Ca_3P_2 is dissolved in water

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78. Complete the reaction :

$P_4 + SOCl_2$ gives

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79. 'X'mol of $CO_2$and 0.05 mol of Ar are enclosed in a vessel of capacity 5 L at 1atm and 27^oC.The value of ' X ' is (R=0.0821 atm $mol^{-1}K^{-1}$)

A. 0.20

B. 0.15

C. .36

D. 0.10

Answer:

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80. What happens when F_2 reacts with water?

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81. 'n'mol of $O_2$and 0.04 mol of Ar are enclosed in a vessel of capacity 5 L at 1atm and 27^oC .The value of ' n ' is (R=0.0821 atm $mol^{-1}K^{-1}$)

A. 0.20

B. 0.15

C. 0.16

D. 0.10

Answer:

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82. 'X'mol of $Na$and 0.06 mol of Cl are enclosed in a vessel of capacity 6 L at 1atm and 27^oC.The value of ' X ' is (R=0.0821 atm $mol^{-1}K^{-1}$)

A. 0.15

B. 0.20

C. 0.18

D. 0.16

Answer:

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83. 'X'mol of $CO$and 0.06 mol of C are enclosed in a vessel of capacity 5 L at 1atm and 27^oC .The value of ' X ' is (R=0.0821 atm $mol^{-1}K^{-1}$)

A. 0.14

B. 0.16

C. 0.15

D. 0.12

Answer:

84. Estimate the number of molecules left in a volume of the size of a pinhead about 1 cubic

mm when the air is pumped out to give a vaccuum of 10^{-6}m Hg at $25^{o}C$.

A. 3.2×10^{7}

B. 2.2×10^{7}

C. 3.2×10^{5}

D. 2.2×10^{5}

Answer:

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85. The average velocity of an ideal gas molecule at 27^oC is 0.5 m/s. The average velocity at 927^oC will be

A. 0.6m/s

B. 1m/s

C. 0.4m/s

D. 1.2m/s

Answer:

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86. The average velocity of an ideal gas molecule at 27^oC is 0.4 m/s. The average velocity at 927^oC will be

A. 0.6m/s

B. 0.8m/s

C. 1m/s

D. 0.4m/s

Answer:

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87. The average velocity of an ideal gas molecule at 27^oC is 0.6m/s. The average velocity at 927^oC will be

A. 1m/s

B. 0.8m/s

C. 1.2m/s

D. 1.4m/s

Answer:

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88. The pH of acid rain is

A. 5.7

B. 5.4

C. 6.5

D. 5.9

Answer:

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89. Reducing smog is the mixture of

A. Smoke, fog and SO_2

B. Smoke, fog and NO_2

C. Smoke, SO_2 and NO_2

D. Fog, SO_2 and NO_2

Answer:

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90. Breakdown of ozone layer by CFCs is mainly due to formation of

A. Fluorine radicals

B. Chlorine radicals

C. Free oxygen atoms (0)

D. Both (1) & (2)

Answer:

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