



CHEMISTRY

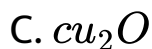
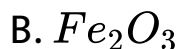
BOOKS - DISHA PUBLICATION CHEMISTRY (HINGLISH)

GENERAL PRINCIPLES AND PROCESSES OF ISOLATION OF ELEMENTS

Jee Main 5 Years At A Glance

1. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of

caprous oxide with



Answer: C

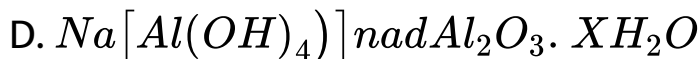
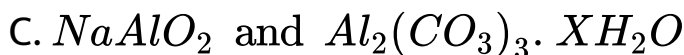


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2. In the leaching method bauxite ore is digested with a concentrated solution of NaOH that produces 'X' whne CO_2 gas is passed through the aqueous

solution of 'X' a hydrated compound 'Y' is precipitated

'X' and 'Y' respectively are:



Answer: D



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3. Extraction of copper by smelting uses silica as an additive to remove:

A. Cu_2O

B. FeS

C. FeO

D. Cu_2S

Answer: C



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4. Which one of the following ores is best concentrated by froth flotation method:

A. Galena

B. Malachite

C. Magnetite

D. Siderite

Answer: A



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5. In the isolation of metals, calcination process usually results in:

A. metal hydroxide

B. metal sulphide

C. metal oxide

D. metal carbonate

Answer: C



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6. Calamine is an ore of

A. zinc

B. aluminium

C. iron

D. copper

Answer: A



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7. In the correct of the Hall-Heroult process for the extraction of Al , which of the following statements is false ?

- A. Al^{3+} is reduced at the cathode to form Al .
- B. Na_3AlF_6 , serves as the electrolyte.
- C. CO and CO_2 are produced in this process.
- D. Al_2O_3 is mixed with CaF_2 which lowers the melting point of the mixture and brings

conductivity.

Answer: B



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8. The form of iron obtained from blast furnace is:

A. Steel

B. Cast Iron

C. Pig Iron

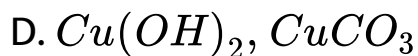
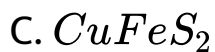
D. Wrought Iron

Answer: B



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9. Which one of the following ores is known as Malachite:



Answer: D



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10. The metal that cannot be obtained by electrolysis of an aqueous solution of its salts is :

A. Ag

B. Ca

C. Cu

D. Cr

Answer: B



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Exercise 1 Concept Builder Topicwise

1. An example of an oxide ore is

A. Bauxite

B. Malachite

C. Zinc blende

D. Feldspar

Answer: A



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2. The most abundant metal in the earth crust is

A. Fe

B. Al

C. Ca

D. Na

Answer: B



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3. Which of the following is an ore of tin ?

A. Carborundum

B. Epsomite

C. Cassiterite

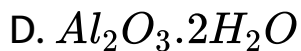
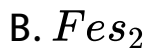
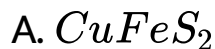
D. Spodumene

Answer: C



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4. Which of the following is chalcopyrite?

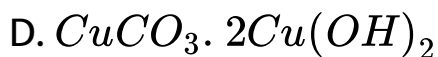
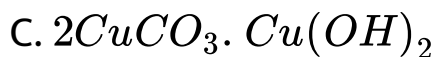
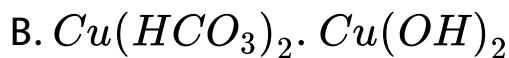


Answer: A



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5. Composition of Azurite mineral of



Answer: C



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6. German silver is an alloy of

A. Ag and Ni

B. Cu, Zn and Ni

C. Au, Cu and Zn

D. Cu and Zn

Answer: B



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7. Steel contains carbon

A. 0.12% to 0.25%

B. 2.5% to 4.5%

C. 1 to 2%

D. 0.5 to 1.5%

Answer: D



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8. The method of concentrating the ore which makes use of the difference in density between ore and impurities is called

A. levigation

B. leaching

C. magnetic separation

D. liquifaction

Answer: A



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9. Magnesium is extracted electrolysing fused magnesium chloride containing $NaCl$ and $CaCl_2$ using:

A. a nickel cathode and a graphite anode.

B. the iron container as anode and a nickel cathode.

C. the iron container as cathode and a graphite anode.

D. the nickel container as cathode and iron anode.

Answer: C



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10. AgCl on fusion with sodium carbonate, gives :

A. Ag_2CO_3

B. Ag_2O

C. Ag

D. Ag_2C_2

Answer: C



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11. Highly electropositive metal(s) can not be commercially extracted by carbon reduction process at high temperature because these :

A. metals combine with carbon to form covalent carbide.

B. metals combine with carbon to form ionic carbide.

C. ΔG_f of highly electropositive metal oxide is having low negative value.

D. metal oxides are not reduced by carbon.

Answer: B



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12. Choose the correct option the code regarding roasting process.

(I) It is the process of heating the ore in air in a

reverberatory furnace to obtain the oxide.

(II) It is an exothermic process.

(III) It is used for the concentration of sulphide ore.

(IV) It removes easily oxidisable volatile impurities present in the concentrated ore.

A. I, II and III

B. I, II and IV

C. I, III and IV

D. I, II, III and IV

Answer: B



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13. Froth floatation process used for the concentration of sulphide ore :

A. is based on the difference in wettability of different minerals.

B. Uses sodium ethyl xanthate, $C_2H_5OS_2$ Na as collector.

C. Uses NaCN as depressant in the mixture of ZnS and PbS when ZnS forms soluble complex and PbS forms froth.

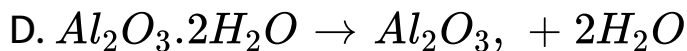
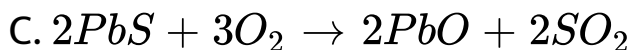
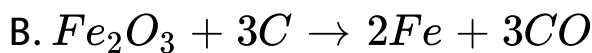
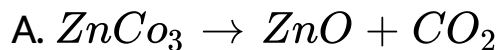
D. All are correct statements.

Answer: D



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14. Which of the following processes involve the roasting process ?



Answer: C



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15. In the froth flotation process for beneficiation of ore, the ore particles float because

- A. They are light
- B. Their surface is not easily wetted by water.
- C. They bear electrostatic charge.
- D. They are insoluble.

Answer: B



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16. Which of the following statement is correct?

- A. Roasting is unnecessarily done for Fe extraction because there is no sulphide ore.
- B. In the smelting step to Cu extraction, reduction of the ore takes place.
- C. Ores may not be mineral.
- D. Sphalerite is the ore of zinc.

Answer: D



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17. Froth floatation process is based on _____.

- A. wetting properties of ore particle.
- B. specific gravity of ore particles.
- C. magnetic properties of ore particles.
- D. electrical properties of ore particles.

Answer: A



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18. Before introducing FeO in blast furnace, it is converted to Fe_2O_3 by roasting so that

- A. it may not be removed as slag with silica.

B. it may not evaporate in the furnace.

C. presence of it may increase the m.pt. of charge.

D. None of these.

Answer: A



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19. Electrolytic reduction method is used for the extraction of

A. highly electronegative elements.

B. highly electropositive elements.

C. metalloids.

D. transition metals.

Answer: B



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20. Calcination is used in matallurgy for removal of

A. moisture

B. water and CO_2

C. CO_2 and H_2S

D. H_2O and H_2S

Answer: B



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21. Heating of pyrites in air for oxidation of sulphur is called

- A. roasting
- B. calcination
- C. smelting
- D. slagging

Answer: A





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22. Which one of the furnaces among the following can produce the highest temperature?

- A. muffle furnace
- B. blast furnace
- C. reverberatory furnace
- D. electric furnace

Answer: D



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23. Furnaces are lined with calcium oxide as

- A. it gives off oxygen on heating.
- B. it gives strong light on heating.
- C. it is refractory and basic.
- D. it is not affected by acids.

Answer: C



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24. Roasting is generally done in case of the

- A. oxide ores

B. silicate ores

C. sulphide ores

D. carbonate ores

Answer: C



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25. Aluminothermic process is used for the extraction of metals, whose oxides are

A. fusible

B. not easily reduced by carbon.

C. not easily reduced by hydrogen.

D. strongly basic.

Answer: B



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26. Frothers (used in Froth floatation process) are

A. molecules with hydrophilic head and hydrophobic tail.

B. molecules with hydrophobic head and hydrophilic tail.

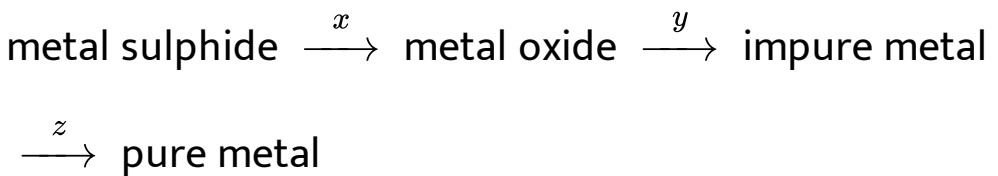
C. both of the above.

D. none of these

Answer: A

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27. Identify x, y, z for the following metallurgical process :



generally x, y and z are respectively:

A. roasting, smelting, electrolysis

B. roasting, calcination, smelting.

C. calcination, auto-reduction, bassemmerisation

D. None of the above is correct

Answer: A



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28. The process of converting hydrated alumina into anhydrous alumina is called

A. roasting

B. smelting

C. dressing

D. calcination

Answer: D



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29. Furnaces are lined with calcium oxide as

A. it gives off oxygen on heating.

B. it gives strong light on heating.

C. it is refractory and basic.

D. it is not affected by acids.

Answer: C



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30. The substance used in the thermite process of reducing metal ores is

- A. aluminium
- B. thorium
- C. heated Pt gauge
- D. carbon

Answer: A



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31. After partial roasting, the sulphide of copper is reduced by

- A. cyanide process
- B. electrolysis
- C. reduction with carbon
- D. self reduction

Answer: D

32. Before introducing FeO in blast furnace, it is converted to Fe_2O_3 by roasting so that

- A. it may not be removed as slag with silica.
- B. it may not evaporate in the furnace.
- C. presence of it may increase the m.pt. of charge.
- D. None of these.

Answer: A

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33. ELLINGHAM DIAGRAM

- A. ΔG increases with an increase in temperature.
- B. It consists of plots of $\Delta_f G^\circ$ Vs T for formation of oxides.
- C. A coupling reaction can be well expressed by this diagram
- D. It express the kinetics of the reduction process.

Answer: D

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34. The form of iron obtained from blast furnace is:

A. Steel

B. Cast Iron

C. Pig Iron

D. Wrought Iron

Answer: B



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35. The slag obtained during the extraction of copper from copper pyrites is composed mainly of

A. $FeSiO_3$



Answer: A



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36. Muffle furnace is used in the metallurgy of

A. Zn

B. Sn

C. Pb

D. Cu

Answer: A



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37. In Hall's process, the main reagent is mixed with

A. NaF

B. Na_3AlF_6

C. AlF_3

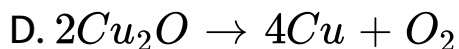
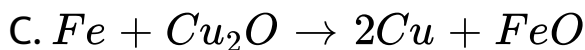
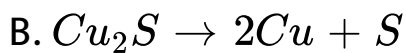
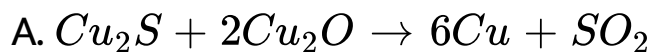
D. none of these

Answer: B



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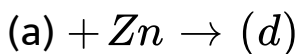
38. In the extraction of copper, the metal formed in the Bessemer converter is due to the reaction



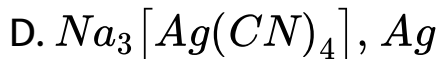
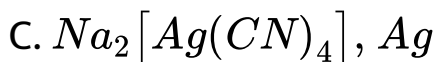
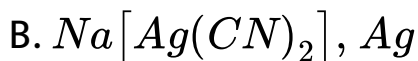
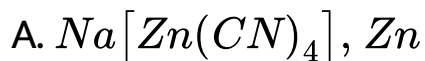
Answer: D



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(b) is a metal. Hence (a) and (b) are



Answer: B



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40. Method used for obtaining highly pure silicon which is used as a semiconductor material, is

- A. oxidation
- B. electrochemical
- C. crystallization
- D. zone refining

Answer: D



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41. In the purification of impure nickel by Mond's process, metal is purified by:

- A. Electrolytic reduction
- B. Vapour phase thermal decomposition
- C. Thermite reduction
- D. Carbon reduction

Answer: B



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42. Incorrect match

A. Purification of Al metal : Baeyer's method

B. Polling: Reduction of Cu_2O

C. $FeCr_2O_4$ (chromite ore) : $NaOH / Na_2CO_3$

D. Ag : Mac Arthur cyanide process

Answer: A



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43. In van Arkel method, if I_2 is introduced at 1700 K over impure metal, the product will be :

A. Iodide of the metal

B. No reaction takes place

C. Impurities react with iodine

D. None of these.

Answer: A



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44. Van Arkel method of purification of metals involves converting the metal to

A. volatile stable compound.

B. volatile unstable compound.

C. non volatile stable compound.

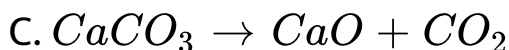
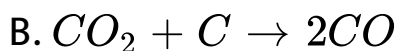
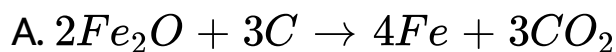
D. None of the above.

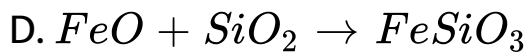
Answer: A



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45. Identify the reaction that does not take place in a blast furnace





Answer: B



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46. The temperature ($^{\circ}C$) at which Fe_2O_3 is finally reduced to Fe in the blast furnace is

A. 993

B. 797

C. 897

D. 1597

Answer: A



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47. High purity copper is obtained by

- A. Carbon reduction
- B. Hydrogen reduction
- C. Electrolytic reduction
- D. Thermite reduction

Answer: C



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48. Iron is obtained on a large scale from Fe_2O_3 by

A. reduction with Na

B. reduction with Al

C. reduction with CO

D. passing H_2

Answer: C



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49. Which of the following metals is obtained by electrolytic reduction process?

A. Fe

B. Cu

C. Ag

D. Al

Answer: D



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50. In the extraction of iron, slag is produced. Slag is

A. CO

B. $FeSiO_3$

C. $MgSiO_3$

D. $CaSiO_3$

Answer: D



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51. Thomas slag is

A. calcium silicate

B. calcium phosphate

C. tricalcium phosphate and calcium silicate

D. calcium ammonium phosphate

Answer: C



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52. Which of the following pairs of metals is purified by van Arkel method?

A. Ga and In

B. Zr and Ti

C. Ag and Au

D. Ni and Fe

Answer: B



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53. Which of the following metals is extracted by the electrometallurgical method?

A. Cu

B. Fe

C. Na

D. Ag

Answer: C



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54. The electrolytic reduction technique is used in the extraction of

- A. highly electronegative elements.
- B. highly electropositive elements.
- C. metalloids
- D. transition metals.

Answer: B





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55. Nickel is purified by thermal decomposition of its

A. hydride

B. chloride

C. azide

D. carbonyl

Answer: D



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56. In the electrolysis of alumina, cryolite is added to

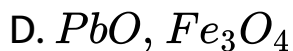
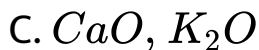
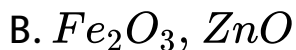
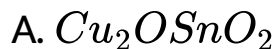
- A. Lower the melting point of alumina and to increase the electrical conductivity
- B. Minimise the anode effect
- C. Remove impurities from alumina
- D. None of these

Answer: A



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57. (viii) Amongst the following groups of oxides, the group containing oxides that cannot be reduced by carbon to give the respective metals is.



Answer: C



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58. The phenomenon of removing layers of basic oxides from metals before electroplating is called

A. galvanising

B. anodising

C. pickling

D. poling

Answer: C



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59. which of the following elements is present as the impurity to the maximum extent in the pig iron?

A. Manganese

B. Carbon

C. Silicon

D. Phosphorus

Answer: B



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60. Carbon and CO gas are used to reduce which of the following pairs of metal oxides for extraction of metals?

A. FeO , SnO

B. SnO , ZnO

C. BaO , Na_2O_2

D. FeO , ZnO

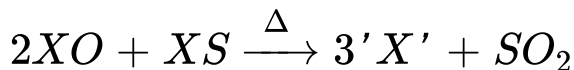
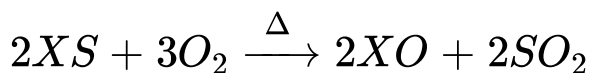
Answer: D



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Exercise 1 Concept Applicator

1. Consider the following reactions :



Then 'X' can not be :

A. Hg

B. Pb

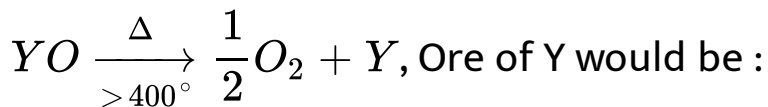
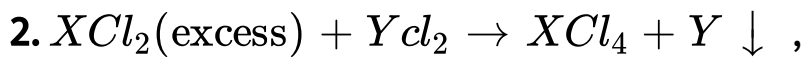
C. Zn

D. Cu

Answer: C



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- A. Siderite
- B. Cinnabar
- C. Malachite
- D. Hornsilver

Answer: B



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3. All ores are minerals while all minerals are not ores because :

A. the metal can't be extracted economically from all the minerals.

B. minerals are complex compounds

C. the minerals are obtained from mines.

D. All of these are correct.

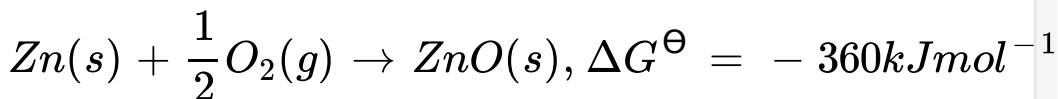
Answer: A



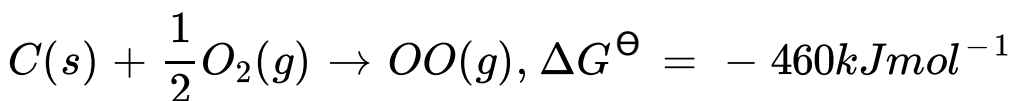
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4. Consider the following reaction at $1000^{\circ}C$

(A)



(B)



Choose the correct statement at $1000^{\circ}C$

- A. ZnO is more stable than CO.
- B. ZnO can be reduced to Zn by C.
- C. ZnO and CO are formed at equal rate.
- D. ZnO can not be reduced to Zn by C

Answer: B



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5. Correct match is:

- (a) Bayer's method - Na_2CO_3
- (b) Matte - 98% Cu_2S + 2% FeS
- (c) van Arkel method - AgI
- (d) Thomas slag - Raw material for cement industry



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6. Select correct statement:

A. The decomposition of an oxide into oxygen and metal vapour entropy increases.

B. Decomposition of an oxide is an endothermic change.

C. To make ΔG° negative, temperature should be high enough so that $T\Delta S^\circ > \Delta H^\circ$

D. All are correct statements.

Answer: D



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7. By which process Pb and Sn are extracted respectively are:

- A. Carbon reduction, self-reduction
- B. Self-reduction, carbon reduction
- C. Electrolytic reduction, cyanide process
- D. Cyanide process, electrolytic reduction

Answer: B



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8. Which of the following facton is of no significance for roasting sulphide ores to the oxide and not subjecting the sulphide ores in carbon reduction directly ?

A. CO_2 is more volatile than CS_2

B. Metal sulphides are thermodynamically more stable than CS_2

C. Metal sulphides are less stable than the corresponding oxides.

D. CO_2 is thermodynamically more stable than CS_2

Answer: B



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9. When copper pyrites is roasted in excess of air, a mixture of $CuO + FeO$ is formed. FeO is present as impurity. This can be removed as slag during reduction of CuO . The flux added to form slag is

- A. SiO_2 , which is an acidic flux
- B. Limestone, which is a basic flux
- C. SiO_2 which is the basic flux
- D. CaO , which is a basic flux

Answer: A



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10. Sulphide ores of metals are usually concentrated by both floatation process. Which of the following sulphide ores offers an exception and is concentrated by chemical leaching ?

A. Galena

B. Copper pyrite

C. Sphalerite

D. Argentite

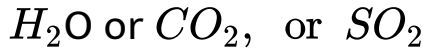
Answer: D



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11. Calcination is the process in which

A. Ore is heated above its melting point to expel



B. Ore is heated below its melting point to expel

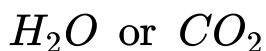
volatile impurities

C. Ore is heated above its melting point to

remove S, As and Sb as SO_2 , As_2O_3 , and Sb_2O_3

, respectively.

D. Ore is heated below its melting point to expel



Answer: D



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12. Extraction of zinc from zinc blende is achieved by

A. roasting followed by self-reduction.

B. electrolytic reduction.

C. roasting followed by reduction with carbon.

D. roasting followed by reduction with another metal.

Answer: C



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13. In froth flotation process many chemicals (frother, collector, activator, and depressant) are used. Which of the following is a frother:

A. $CuSO_4$

B. $NaCN$ + alkali

C. Pineoil

D. Potassium xanthate

Answer: C



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14. In the metallurgy of Zn the Zn dust obtained from roasting and reduction of zinc sulphide contains some ZnO. It is removed by

A. absorbance of ultraviolet light and reemission of white light.

B. shock cooling by contact with a shower of molten lead.

C. X-ray method.

D. smelting

Answer: D



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15. A sulphide ore is first converted into its oxide before reduction. This is done because :

A. the enthalpy of formation of CO_2 is more than that of CS_2 .

B. a metal sulphide is generally more stable than the metal oxide.

C. no reducing agent is found suitable for reducing a sulphide ore.

D. a sulphide ore cannot be reduced at all.

Answer: A



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16. Carbon and CO gas are used to reduce which of the following pairs of metal oxides for extraction of metals?

A. FeO, SnO

B. SnO, ZnO

C. BaO, Na_2O_2

D. FeO , ZnO

Answer: D



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17. While extracting an element from its ore, the ore is grind and leached with dil KCN solution to form the soluble product potassium argentocyanide. The element is

A. Lead

B. Chromium

C. Manganese

D. Silver

Answer: D



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18. Consider the following statements -

In the aluminothermite process, aluminium acts as reducing agent.

The process of extraction of gold involves the formation of $[Au(CN)_2]^-$ and $[Zn(CN)_4]^{2-}$

In the extractive metallurgy of zinc, partial fusion of ZnO with coke is called sintering and reduction of ore to the molten metal is called smelting.

Extractive metallurgy of silver from its ore argentite involves complex formation and displacement by

more electropositive metal. Choose the correct options -

A. A and B

B. B and C

C. A, B and C

D. A, B, C and D

Answer: D



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19. Which statement is correct?

A. Gangues are carefully chosen to combine with the slag present in the ore to produce easily fusible flux to carry away the impurities.

B. Slags are carefully chosen to combine with the flux present in the ore to produce easily fusible gangue to carry away the impurities.

C. Gangues are carefully chosen to combine with the flux present in the ore to produce easily fusible slag to carry away the impurities.

D. Fluxes are carefully chosen to combine with the gangue present in the ore to produce easily

fusible slag to carry away the impurities.

Answer: D



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20. Aluminium is prepared in large quantities by

A. heating cryolite in a limited quantity of air

B. reducing aluminium oxide with coke.

C. reducing aluminium oxide with sodium.

D. electrolysing aluminium oxide dissolved in
fused electrolyte

Answer: D



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21. The incorrect statement among the following is

A. Hydrogen is used to reduce NiO

B. Zirconium is refined by Van Arkel method.

C. The sulphide ore galena is concentrated by froth flotation.

D. In the metallurgy of iron, flux used is SiO_2 –

Answer: D

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22. Select the property which forms the basis of hydrometallurgical process that is used for extraction of gold.

- A. Gold is electropositive metal
- B. Gold is less reactive metal
- C. Gold forms complexes that are water soluble.
- D. Gold forms salts, which are water soluble.

Answer: C

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23. which of the following elements is present as the impurity to the maximum extent in the pig iron?

A. Manganese

B. Carbon

C. Silicon

D. Phosphorus

Answer: B



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24. Which of the following pairs of metals is purified by van Arkel method?

A. Ga and In

B. Zr and Ti

C. Ag and Au

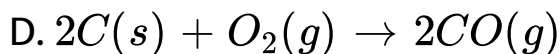
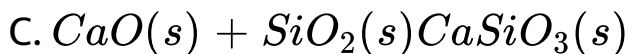
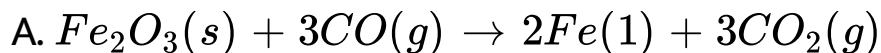
D. Ni and Fe

Answer: B



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25. The following reaction take place in the blast in the preparation of impure iron identify the reaction pertaining to the formation of the slag

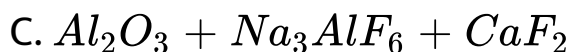
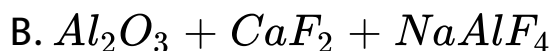


Answer: C



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26. Aluminium is extracted from Alumina (Al_2O_3) by electrolysis of a molten mixture of



Answer: C



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27. Which ore of the following is a mineral of iron ?

A. Malachite

B. Cassiterite

C. Pyrolusite

D. Magnetite

Answer: D



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28. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

A. Copper (I) sulphide (Cu_2S)

B. Sulphur dioxide (SO_2)

C. Iron sulphide (FeS)

D. Carbon monoxide (CO)

Answer: A



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29. 'Metals are usually not found as nitrates in their ores'. Out of the following two (*I* and *II*) reasons which is//are true for the above observation?

I. Metal nitrates are highly unstable.

II. Metal nitrates are highly soluble in water.

- A. (i) and (ii) are false
- B. (i) is false but (ii) is true
- C. (i) is true but (ii) is false
- D. (i) and (ii) are true

Answer: B



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30. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

- A. Iron (II) sulphide

B. Carbon monoxide

C. Copper (I) sulphide

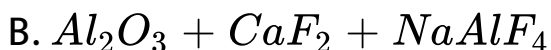
D. Sulphur dioxide

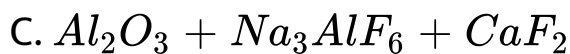
Answer: A



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31. Aluminium is extracted from Alumina (Al_2O_3) by electrolysis of a molten mixture of





Answer: C



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D. Magnetite

Answer: D



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33. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

- A. Copper (I) sulphide (Cu_2S)
- B. Sulphur dioxide (SO_2)
- C. Iron sulphide (FeS)
- D. Carbon monoxide (CO)

Answer: A



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34. 'Metals are usually not found as nitrates in their ores'. Out of the following two (*I* and *II*) reasons which is//are true for the above observation?

I. Metal nitrates are highly unstable.

II. Metal nitrates are highly soluble in water.

A. (i) and (ii) are false

B. (i) is false but (ii) is true

C. (i) is true but (ii) is false

D. (i) and (ii) are true

Answer: B



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35. In the extraction of copper from its sulphide ore, the metal is finally obtained by the reduction of cuprous oxide with

- A. Iron (II) sulphide
- B. Carbon monoxide
- C. Copper (I) sulphide
- D. Sulphur dioxide

Answer: A



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