



CHEMISTRY

BOOKS - OSWAAL PUBLICATION CHEMISTRY (KANNADA ENGLISH)

Sample Paper 5

Exercise

1. Why are transition metal compounds coloured ?



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2. What happens to change in entropy when water is converted into ice ?



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3. What is metallic bond?



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4. How many moles of $AgCl$ will be precipitated when an excess of $AgNO_3$ solution is added to one molar solution of $[CrCl(H_2O)_5]Cl_2$?



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5. What is secondary cell?



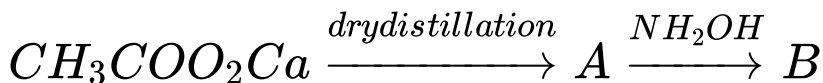
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6. Give the IUPAC name of $[Zn_2Fe(CN)_6]$



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7. Name the end product (B) of the following reaction:



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8. What is the reducing agent used in Clemmensen's reduction?



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9. Give an example for a disaccharide that contains fructose units in furanose form .



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10. How many α - amino acid units are presents in a molecule of insulin ?



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11. What is an Electrolytic cell?



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12. State and derive Ostwaid dilution law for a weak electrolyte.



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13. Write any two differences between order and molecularity of a reaction?



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14. What is lanthanide contraction? Write any one consequence of lanthanid contraction.



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15. What are vinyl and allyle alcohols? Give one example of each.



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16. How test benzaldehyde react with a concentrated solution of sodium hydroxide? Give the equation and name the reaction.



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17. What are antioxidants? Give an example.



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18. What is a broad spectrum antibiotics? Give an example.



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19. Describe parke's process of desilverisation of lead.



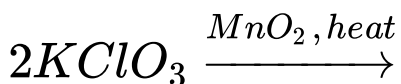
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20. Explain the manufacture of ammonia by Haber's process.



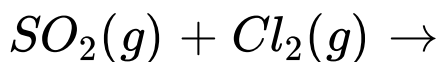
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21. Complete the following equations:



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22. Complete the following equations:



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23. Complete the following equations:



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24. Zinc liberates hydrogen gas when reacted with dilute nitric acid whereas copper does not. Why ?



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25. How concentrated sulphuric acid reacts with potassium bromide.



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26. Explain colour of transition metal compounds based on crystal field splitting.



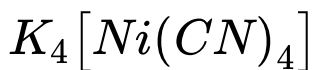
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27. Describe experiment to determine the mass of $KMnO_4$ present in $1dm^3$ of the solution using standard ferrous ammonium sulphate solution.



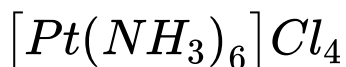
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28. Write the IUPAC names of the following:



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29. Write the IUPAC Name of the



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30. Give an example for a cationic complex.





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31. With the help of valence bond theory explain hybridization, geometry and magnetic property of Nickel tetracarbonyl.



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32. Derive an intergrated rate for the first order reaction.



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33. On the basis of electron gas theory explain the metallic lusture property.



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34. Define relative lowering of vapour pressure.



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35. Write any two demerits of standard hydrogen electrode.



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36. Explain construction of Daniel cell. Write half cell reactions. How is Daniel cell symbolically represented?



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37. Write the equations of anodic and cathodic reactions occur during rusting of iron.



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38. Derive an expression for the density of solids.



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39. Define Energy of Activation.



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40. Write any two differences between physisorption and chemisorption.



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41. Name the phenomenon/effect for the following :

Colloidal particles are in zig-zag motion



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42. Name the phenomenon/effect for the following :

When an electrical potential is applied across two platinum electrodes dipping in colloidal solution, particles moves towards one or the other electrodes



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43. Name the phenomenon/effect for the following :

Scattering of light by colloidal sol



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44. How is p-bromo acetanilide is prepared in the laboratory from acetanilide?



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45. Name the organic compound obtained when aniline is added to dilute HCl at $0^{\circ}C$.
Give the equation for the reaction .



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46. How do you convert ethanoic acid to methane? Write equation.



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47. Explain Clemmensen reduction with an example.



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48. Give the IUPAC name of acetic acid.



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49. Explain the preparation of ethene from ethanol.



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50. Explain Williamson's ether synthesis.



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51. Mention a general test for the following:

Carbohydrates



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52. Mention a general test for the following:

Oils and fats



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53. Give four biological importance of carbohydrates.



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54. When glucose is heated with red phosphorous and hydroiodic acid it forms hexane. What conclusion is drawn from this observation regarding the structure of glucose?



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55. Explain the preparation of Buna-N.



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56. Give an example for thermosetting polymer.

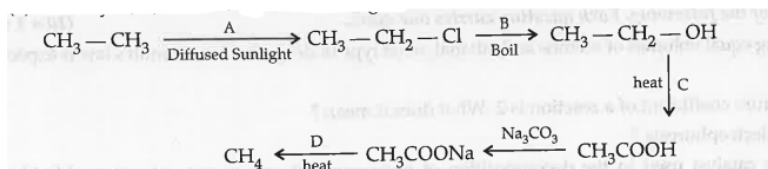


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57. Name the monomers used in the preparation of polythene and natural rubber.

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58. Identify A, B, C and D in the following reactions:



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59. How is ethyl bromide converted into ethylbenzene by Friedel - Crafts reactions ?

Give the equations.



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60. $CH_3Br + AgF \rightarrow CH_3F + AgBr$. Name the reaction.



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