



CHEMISTRY

BOOKS - OSWAAL PUBLICATION CHEMISTRY (KANNADA ENGLISH)

Sample Paper 8

Excercise

1. Give an example of polydentate ligand.



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2. Which allotrope of sulphur is thermally stable at room temperature?



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3. Mention the hybridised state of carbonyl carbon atom.



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4. Give an example for zero order reaction.



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5. Name the method used for isolation of noble gas mixture from air.



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6. What is the gas liberated at anode during the manufacture of caustic soda using Nelson's cell?



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7. How many Faradays are required to liberate 224 cm^3 of H_2 from acidified water at STP?



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8. Write the molecular formula of chromyl chloride.



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9. Name the dispersion medium in gel.



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10. Name the organic compound formed when chlorobenzene is treated with sodium in dry ether.



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11. Aluminium crystallizes in an fcc structure. Atomic radius of the metal is 125 pm. Calculate the edge length of the unit cell of the metal.



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12. Write Gibb's equation



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13. What are two criteria for effective collision according to collision theory?



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14. Give reasons:

Actinoids show variable oxidation states.



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15. Give reasons:

Zr and Hf have almost identical radii.



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16. Give two examples for food preseervaties.



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17. Name the product formed when phenol is treated with acidified solution of $Na_2Cr_2O_7$

Give equation.



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18. Calculate the magnetic moment of Mn^{2+} ion.

[Atomic number of Mn=25]



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19. How does phenol react with bromine water?



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20. (i) Draw a neat labelled diagram of blast furnace used for the extraction of iron and mention different zones.

(ii) What is the role of coke and lime stone during the extraction of cast iron?



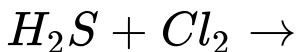
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21. Complete the following equation:



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22. Complete the following equation:



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23. Complete the following equation:





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24. White phosphorous is heated with excess of dry chlorine to get X. X on hydrolysis finally forms an oxoacid of phosphorous Y. What are X and Y?

What is the basicity of the acid ?



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25. Give two reasons why transition elements can form complexes.



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26. In the coordination compound potassium ferrocyanide identify the ligand, central metal ion, primary valency and co-ordination number of the central metal.



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27. Write the IUPAC name of $(\text{CO}(\text{ONO})(\text{NH}_3)_5)\text{Cl}_2$.



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28. State EAN rule for Co-ordination compounds. How many ions are formed the complex, $[\text{Co}(\text{NH}_3)_6]\text{Cl}_2$ in solution.



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29. What is a binary solution? Give an example.



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30. Calculate the number of particles per unit cell in fcc.



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31. What is osmotic pressure?



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32. State Van't Hoff-Boyle's law.



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33. What are isotonic solutions?



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34. Calculate the electrode potential developed when a silver electrode is dipped in 0.025 M silver nitrate solution at 289K.



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35. Define energy of activation. Draw a diagram of energy profile to show the influence of a positive catalyst on the energy of activation of a reaction.



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36. Describe an experiment to show the effect of concentration on the rate of the reaction between potassium persulphate and potassium iodide.



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37. Rate constant of a first order reaction A products is 0.016 min^{-1} . Calculate the time required for 80% of the reaction to be completed.



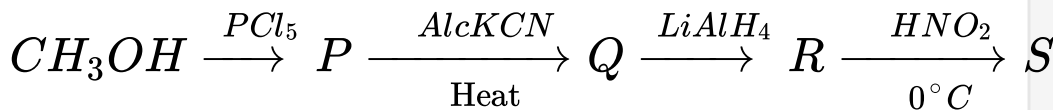
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38. What is peptising agent? How is ferric hydroxide sol prepared?



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39. Identify P,Q,R and S in the following conversions:



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40. Explain Kolbe's reaction with an example.



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41. Name a metal used as catalyst for dehydrogenation of primary alcohol.



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42. Explain Cannizzaro's reaction taking benzaldehyde as an example.



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43. Write the general equations for the reaction of an aldehyde with (i) HCN (ii) Hydrazine.



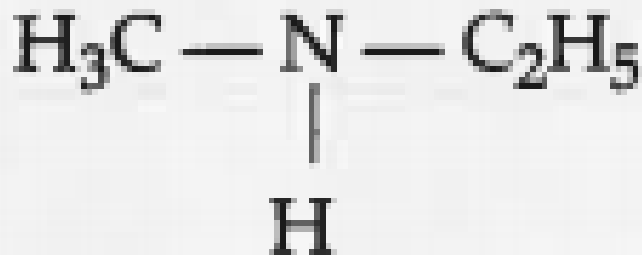
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44. Explain Hoffmann bromamide degradation for the preparation of aniline.



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45. Give IUPAC name of:



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46. b) What is Hinsberg's reagent? Between CH_3NH_2 and $C_6H_5NH_2$ which is more basic?



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47. Why is benzene stable?



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48. How are following formed?

(i) Cinnamic acid from benzaldehyde.

(ii) Acetyl chloride from acetic acid.



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49. How are following formed?

(i) Cinnamic acid form benzaldehyde.

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50. (a) Write the Harworth's structure of maltose.

(b) What are hormones? Give one biological

function of insulin.

(c)What are nucleosides?



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51. Explain refining of oils.



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52. How many secondary-OH groups are present in a molecule of glucose?



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53. What is a thermoplastic? Name the monomer of nylon-6 and give the partial structure of nylon-6.



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54. What is biodegradable polymer ? Give an example of a biodegradable aliphatic polyester.



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