



CHEMISTRY

BOOKS - SHARAM PUBLICATION

SET 18

Exercise

1. In calcium fluoride crystal, which ion occupies the tetrahedral void?



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2. Which organic compound is formed when methyl amine is treated with chloroform and aq.KOH solution?



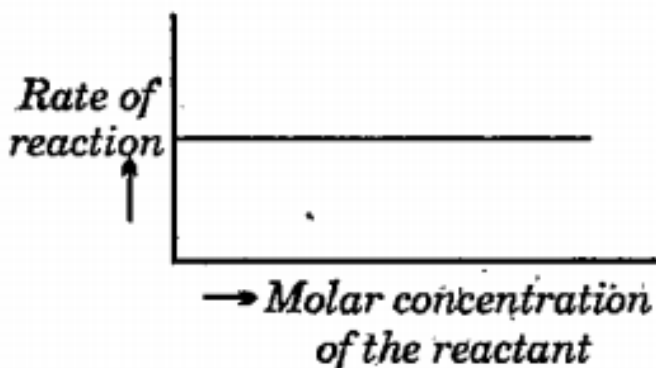
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3. Write the name of one ore of iron and aluminium.



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4. The following graph is obtained when we plot rate of reaction versus molar concentration of the reactant. What is the order of the reaction that is evident from the graph?



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5. Give two examples of antifertility drugs.



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6. Which type of solid is wax?



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7. What happens when sodium benzoate is heated with soalime ?



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8. Chloroprene is the monomer unit of

A. Polyisoprene

B. Neoprene

C. PVC

D. Polythene

Answer:



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9. $FeCl_3$ Solution stops bleeding because

- A. The blood is coagulated and blood cells are sealed
- B. It seals the blood cells
- C. Blood starts flowing in reverse direction
- D. All of these

Answer:



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10. The impurities present in the ore are removed by adding

A. Flux

B. Slag

C. Gangue

D. Another Ore

Answer:



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11. The crystal lattice of CsCl is:

- A. face-centred
- B. Body-centred
- C. Simple cubic
- D. Hexogonal close packing

Answer:



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12. Which of the following are used as analgesic?

A. Aspirin

B. Heroin

C. Pomethazine

D. Serotonin

Answer:



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13. Which of the following does not form pentahalide ?

A. Polystyrene

B. As

C. Sb

D. N

Answer:



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14. When a mixture of calcium benzoate and calcium acetate is dry distilled ,the resulting compound is

- A. acetophenone
- B. benzophenone
- C. benzaldehyde
- D. acetone

Answer:



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15. Define coagulation



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16. Write the structure of two oxoacids of halogen.



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17. What is the difference between lyophobic and lyophilic colloids.



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18. Define Hardy -Schulze rule.



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19. Why do transition elements show variable oxidation states? Write the name of a 3d-series element that exhibits maximum oxidation states.



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20. What is a Galvanic Cell? Write the cell reaction of the Galvanic Cell.



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21. What happens when SO_2 gas is passed through acidified $K_2Cr_2O_7$ solution ?



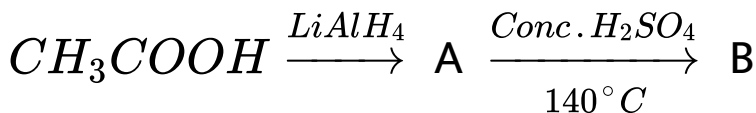
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22. What is the reaction of Mg_3N_2 with water?



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23. Identify the compounds (A) and (B) in the reaction given below.



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24. Give a chemical test to distinguish between ethyl alcohol and methyl alcohol.



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25. Discuss the general trend in properties of transition elements with reference to ionization energy.



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26. Differentiate between multimolecular and macromolecular colloids.



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27. How catalyst affects the rate of reaction ?



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28. Discuss S_N2 reaction.



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29. Describe a method for the identification of primary, secondary and tertiary amines



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30. Write notes on

Brownian movement



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31. What is crystal field splitting among octahedral and tetrahedral crystal fields? In which case, the magnitude of crystal field splitting is larger?



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32. Write a note on interstitial compounds.



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33. Write a note on aldol condensation.



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34. What is Fittig and Wurtz-Fittig reaction?



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35. How is ethyl methyl ether prepared ? Write the reaction of diethyl ether with





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36. How is ethyl methyl ether prepared ? Write the reaction of diethyl ether with

Conc. H_2SO_4



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37. Why is dimethyl ether less soluble in water than ethyl alcohol?



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38. Give a method of preparation of formaldehyde from methyl bromide.



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39. What happens when formaldehyde of treated with Tollen's reagent



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40. What happens when formaldehyde reacts with concentrated $NaOH$ solution?



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41. How is HCl prepared from $NaCl$ in the laboratory?



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42. What happens when HCl gas is passed through

acidified $K_2Cr_2O_7$ solution



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43. What happens when HCl gas is passed through

Leadnitrate solution



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44. What is Aquaregia?



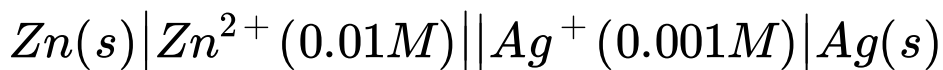
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45. Give a comparison between electrolytic cell and electrochemical cell ?



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46. Calculate the emf of the following cell.



Given $E_{\text{Zn} | \text{Zn}^{2+}}^{\circ} = 0.76V$ and

$$E_{\text{Ag} | \text{Ag}^+}^{\circ} = -0.80V$$



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47. What is a dry cell ? Give an example.



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48. From Arrhenius equation derive the

relation $\log\left(\frac{k_2}{k_1}\right) = \frac{Ea}{2.303R} \left(\frac{T_2 - T_1}{T_1T_2}\right)$

where the symbols have their usual meanings.



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49. The rate constant for a reaction of zero order of A is $0.0030 \text{ mol L}^{-1} \text{ s}^{-1}$. How long will it take for the initial concentration of A to fall from 0.10 M to 0.075 M?



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50. Write the important characteristics of zero order reaction.



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