CHEMISTRY

BOOKS - MBD CHEMISTRY (ODIA ENGLISH)

CLASSIFICATION OF ELEMENTS AND PERIODICITY IN PROPERTIES

Question Bank

1. Which of the following has the largest size ?

A. Be

B. I

C. I^-

D. F

Answer:

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2. To which group of the periodic table does the element havig atomic number 19 belong ?

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3. Which of the following has the smallest size ?

A. Na^+

B. Mg^{2+}

C. Al^{3+}

D.

Answer:

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4. Name any two elements of Group-4 of the periodic table.

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5. Name any two elements of Group vA of periodic table.

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6. Name the element that are included in the nitrogen family.

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7. Name any two elements of IVA of periodic table.

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8. Name any two elements of group VIA of periodic table.

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9. Out of the atomic numbers given below suggest which one of the following will be d-block element and name the element. (Atomic number 12, 13, 18 and 28)

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10. Which of the following is a 'd' block element ?

A. Na

B. Fe

C. Mg

D. C

Answer:

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11. Which of the following has largest size ?

A. Na

B. Na^+

C. N

D. F

Answer:

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12. Between Li and Na which is more electropositive ?

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13. Which of the following has the lowest electron affinity ?

A. K

B. B

C. C

D. F

Answer:

14. Name any two elements of Group IIA of the periodic table.

15. Name an element which can belong bith to group IA and group VIIA of the periodic table.

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16. Which of he following is a 'p'-block element in the periodic table ?

A. Na

B. Mg

C. Ni

D. Mn

Answer:

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17. Name the elements of group IA of the periodic table .

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18. Which of the following has the lowest atomic volume ?

A. Al

B. B

C. F

D. O

Answer:

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19. Which is the most electronegative element in the periodic table ?

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20. What happens to the atomic size of the element along a period ?

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21. Which of the following elements has the smallest atomic size ?

A. C

B. F

C. Na

D. Cl

Answer:

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22. Which of the following elements has the highest electronegativity ?

A. F

B. He

C. Na

D. Ca

Answer:

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23. Which of the following is a 'p-' block element in the periodic table ?

A. Na

B. O

C. Mg

D. Fe

Answer:

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24. How atomic radii vary in a group of periodic table ?

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25. How does ionisation energy vary in a group ?

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26. How does the electron affinity of the elements present in a group change ?

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27. From amongst the following elements choose one p-block element ?

A. K

B. N

C. Cr

D. Ca

Answer:

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28. Which of the following is a d-block element in the periodic table ?

A. K

B. N

C. Cr

D. Ca

Answer:

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29. Name an element which can belong bith to group IA and group VIIA of the periodic table.

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30. Name any two element from d-block in long form of periodic table .

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31. Which of the following elements has the highest ionisation potential ?

A. Na

B. Ne

C. Ca

D. Mg

Answer:

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32. Whiuch of the following elements has the lowest electronegativity ?

A. P

B. N

C. O

D. F

Answer:

Watch Video Solution

33. Which of the following has the smallest size ?

A. Na^+

B. Mg^2+

C. Al^3+

D.

Answer:

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34. What is the electronic configuration of the element of atomic number 22. To which group of the periodic table does it belong ?

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35. State Mendeleev's periodic law.

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36. Electron affinity of flourine is less than that of chlorine Why ?

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37. Atomic volume is not the fundamental property of an atom. Explain, in five sentences.

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38. Write short note on characterstics of d-block elements.

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39. The ionic radii of alkaline earth metals are smaller than those of the nerest alkali metals. Why ?

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40. What do you understand by electron affinity ? How electron affinity vary in a periodic table ?

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41. Explain with reason : Alkali metals do not form dipositive ions.

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42. Expalin with reason : Electronegativity values of inert gases are zero.

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43. The electronic configuration of an element in $1s^2, 2s^2, 2p^6, 3s^2, 3p^6$. In which group of the

periodic table will it be placed and why ?

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44. Why electron affinity of nitrogen is zero.

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45. Name the elements of group IA of the periodic table .

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46. Describe in brief the main feature of long form of periodic table.What are s-block elements ?

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47. What is the long from periodic table ? What are the advantages of this peroidic table ?

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48. What is periodic law ? What advantages the long form of periodic table has got over Mendeleev's periodic table ?

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49. How the properties of the elements vary in each group and each period of the periodic table ?

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50. What is the long from periodic table ? What are the advantages of this peroidic table ?

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51. What are electron affinity and ionisation energy ? Expalin, Arrange nitrogen, oxygen and flourine in the ncresing order of their ionisation energies giving reasons.

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52. What is long form periodic table ? How do the properties of elements vary in groups and periods of the periodic table? Name the elements of zero group.

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53. Describe long form of periodic table . How ionisation energy vary in groups and periods ?

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54. The outermost elecronic configuration of most electronegative element is :

A. ns^2np^3

B. ns^2np^4

C. ns^2np^5

D. ns^2np^6

Answer: C

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55. The element colifornium belongs to the family of :

A. actinide series

B. alkaline -earth family

C. lanthanide series

D. alkali metal family

Answer: A

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56. Which of the followig structure is associated with the biggest jump between the second and the third ionisation energies ?

A. $1s^2, 2s^2, 2p^6$

B. $1s^2, 2s^2, 2p^6, 3s^1$

C. $1s^2, 2s^2, 2p^6, 3s^2$

D. $1s^2, 2s^2, 2p^1$

Answer: C

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57. Which of the follwing represents the correct order of electeon affinities ?

A. Cl gt Br gt I gt F

B. Cl gt F lt Br lt I

C. F gt Cl gt Br gt I

D. Cl gt F gt Br gt I

Answer: D

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58. If the atomic number of an element is 33, it will be placed in Mendeleev's periodic table in the :

A. first group

B. third group

C. fifth group

D. seventh group

Answer: C

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59. Which one of the following ions has the smallest radius ?

A. Cl^-

B. S^{2-}

C. K^+

D. Ca^{2+}

Answer: D

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60. Which of the followig structure is associated with the biggest jump between the second and the third ionisation energies ?

A. $1s^2, 2s^2, 2p^6, 3s^1$

B. $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$

C. $1s^2, 2s^2, 2p^6, 3s^2, 3p^2$

D. $1s^2, 2s^2, 2p^6, 3s^2$

Answer: D

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61. Which of the following is incorrect with respect to the property indicated ?

A. Electronegativity: $F_2 > Cl_2 > Br_2$

B. Electron affinity : $Cl_2 > F_2 > Br_2$

C. Oxidation power: $F_2 > Cl_2 > Br_2$

D. Bond energy : $F_2 > Cl_2 > Br_2$

Answer: D

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62. The first ionosation potential will be maximum for :

A. uranium

B. iron

C. hydrogen

D. lithium

Answer: C

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63. The correct order of first ionisation potential among following elements , Be, B, C, N and O is :

A. B ltBe lt C lt O lt N

B. B ltBe lt C lt N lt O

C. B ltBe lt C lt N lt O

D. Be ltB lt C lt O lt N

Answer: A

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64. General electronic configuration of lanthanide is :

A. $(n-2)f^{1-14}, (n-1)s^2p^6d^{0-1}, ns^2$

B. $(n-2)f^{10-14}, (n-1)d^{10-1}, ns^2$

C. $(n-2)f^{0-14}, (n-1)d^{10}, ns^2$

D. $(n-2)d^{0-1}, (n-1)f^{1-14}, ns^2$

Answer: A

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65. Element with atomic number 56 belongs to which block ?

A. s

B. p

C. d

D. f

Answer: D

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66. The ions $O^{2-}, Na^{+}, F^{-}, Mg^{2+}$ and Al^{3+} are isoelectronic,Their ionic radii show

A. a significant decrease from $O^{2-} \rightarrow Al^{3+}$

B. an increase from $O^{2-} \rightarrow F^{-}$ and then decrease from $Na^{+} \rightarrow Al^{3+}$

C. a decrease from O^{2-} to F^{-} and then increase from Na^{+} to Al^{3+}

D.

Answer: A

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67. The atomic number if vanadium (V), chromiun (Cr), manganese (Mn) and iron (Fe) are respectively 23, 24 , 25, and 26 Which one of these may be expected pt have the highest second ionization enthalphy ?

A. Fe

B. V

C. Cr

D. Mn

Answer: A

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68. Which one of the following groupinga represents a collection of isoelectronic species ?

[At.nos- Cs - 55 , Br - 35]

A. Ca^{2+} , Cs^{+} , Br

B. Na^+, Ca^{2+}, Mg^{2+}

C. N^{3-}, F^-, Na^+

D. Be, Al^{3+}, Cl^-

Answer: C

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69. What is the characteristic valence shell configuration of coinage metals ?

A. ns^2np^6

B. $(n-1)d^2np^6$

C. nd^9ns^1

D. $(n-1)d^{10}ns^1$

Answer: D

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70. Which is an amphoteric oxide ?

A. ZnO

B. CaO

C. BaO

D. SrO

Answer: A

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71. The order of E.A of F, Cl, Br, and I is :

A. F lt Clgt Brgt I

B. F gt Clgt Brgt I

C. F lt Cl lt Br lt I

D. F gt Cl lt Br lt I

Answer: A

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72. Which one of the following sets of ions represents a collection of iso-electronic species ?

A. $k^+, Cl^-, Mg^{2+}, Sc^{3+}$

B. $Na^+, Ca^{2+}, Sc^{3+}, F^-$

C. $K^+, Ca^{2+}, Sc^{3+}, CL^-$

D. $Na^+, Mg^{2+}, Al^{3+}, Cl^-$

Answer: C

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73. Which of the following electronic comfigurations has the lowest ionosation energy ?

A. $1s^2 2s^2$ 2p^6`

B. $1s^2 2s^2 2p^3$

C. $1s^2 2s^2 2p^6 3s^1$

D.

Answer: C

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74. Name the elements of group IA of the periodic table .

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75. Which two element of the second period of the periodic table have both positive and negative oxidation number ?

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76. Which of the following has the smallest size ?

A. Na^+

B. Mg^(2+)`

C. AL^{3+}

D.

Answer: C

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77. Which halogen shows only one oxidation number in its compounds ?

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78. Whiuch of the following elements has the lowest electronegativity ?

A. P

B. N

C. O

D. F

Answer: A

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79. Which of the following elements has the highest ionisation potential ?

A. Na

B. He

C. Ca

D. Mg

Answer: B

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80. Arrange the following in order of increasing electronegativity: S, CL, F, Cs

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81. Arrange the following in order of increasing size: Cl, Cl^+, Cl^-

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82. Name any two element from d-block in long form of periodic table .

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83. Which of the following has largest size ?

A. Na

B. Na^+

C. N

D. F

Answer: A

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84. Which of the following has the small size ?

A. Cl^-

B. Cl

C. O^{2-}

D.

Answer: C

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85. Which of the following is a d-block element in the periodic table ?

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86. Name an element which can belong bith to group IA and group VIIA of the periodic table.

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87. How does the electron affinity of the elements present in a group change ?

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88. From among the following elements choose one p-block element:

A. Na

B. Be

C. Ca

D. Al

Answer: D

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89. Which of the following is a 'p-' block element in the periodic table ?

A. Na

B. O

C. Mg

D. Fe

Answer: B

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90. How do atomic radii vary in a group of periodic table ?

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91. How does ionisation energy vary in a group ?

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92. From amongst the following electronic configuration, find out the alkaline earth element.

$[Ar]4s^2$

$[Ar]3d^5 4s^2$

$[Ar]3d^{10} 4s^2$

A. [Ar] $4s^2$,

B. [Ar] $3d^{54}s^2$

C. [Ar]$3d^{10}4s^2$

D.

Answer: A

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93. What happens to the atomic size of the element along a period ?

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94. Which of the following elements has the smallest atomic size :

A. Cs

B. F

C. Na

D. Cl

Answer: B

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95. Which of the following elements has the highest electronegativity ?

A. F

B. He

C. Na

D. Ca

Answer: A

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96. Which is the most electronegative element in the periodic table ?

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97. How atomic radii vary in a group of periodic table ?

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98. Name any two elements of Group IIA of the periodic table.

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99. Which of he following is a 'p'-block element in the periodic table ?

A. Na

B. Mg

C. N

D. Mn

Answer: C

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100. Name any two elements of group IIA of periodic table.

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101. Which of the following has the lowest atomic volume ?

A. Al

B. B

C. F

D. O

Answer: C

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102. Which of the following is a 'd' block element ?

A. Na

B. Fe

C. Mg

D. C

Answer: B

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103. Which of the following has the largest size ?

A. Na

B. Na^+

C. N

D. F.

Answer: A

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104. Write the electronic configuration of second element of group VI-A of the periodic table.

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105. Between lithium and sodium which is more electropositive ?

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106. Which of the following has the lowest electron affinity ?

A. K

B. B

C. C

D. F

Answer: A

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107. Name two elements of Group VA of periodic table.

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108. Out of the atomic numbers given below suggest which one of the following will be a d-group element and name the element.

A. 12

B. 13

C. 18

D. 25

Answer: D

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109. Name the element present in group IIA of thethird period of the periodic table.

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110. Which of the following has largest size ?

A. Na

B. Na^+

C. Cl

D. Cl^-

Answer: D

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111. Name any two elements of IVA of periodic table.

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112. Name any two elements of IVA of periodic table.

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113. Which of the following has the smallest size ?

A. Na^+

B. Mg^{2+}

C. Al^{3+}

D.

Answer: C

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114. How does electronegativity vary in a group and period ?

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115. What is oxidation number ?

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116. How many atoms of fluorine are there in 1.9×10^{-6} gms of fluroine ? (F = 19 amu)

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117. Why Mg^{2+} ion is smaller than O^{2-} althrough both have the same electronic

configuration ?

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118. Why ionosation energy of magnesium is greater than aluminium ?

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119. Why electron affinity of Be is zero ?

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120. What are sigma and pi-bond's ? What is the bond angle and shape of the methane molecule ?

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121. Why electron affinity of nitrogen is zero.

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122. What is the sum of the oxidation number of all the atoms in a neutral molecule ?

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123. Explain with reason : Alkali metals do not form dipositive ions.

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124. Expalin with reason : Electronegativity values of inert gases are zero.

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125. Justify the position of carbon and lead in the periodic table on the basis of electronic configuration.

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126. The electronic configuration of an element in $1s^2, 2s^2, 2p^6, 3s^2, 3p^6$. In which group of the periodic table will it be placed and why ?

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127. Explain on the basis of electronic configuration why fluorine and chlorine are put in VIIA of the periodic table.

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128. The electronic configuration of an element in $1s^2, 2s^2, 2p^6, 3s^2, 3p^6$. In which group of the periodic table will it be placed and why ?

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129. Among the four halogens name one which has :

highest value of electron affinity?

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130. What is oxidation number ?

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131. Explain why the ionisation energy of nitrogen is more than that of oxygen.

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132. The atomic numbers of N,Be and Mg are 7, 4 and 12 respectively. Explain why the electron affinites of hese elements are almost zero.

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133. Electron affinities of halogens are high why ?

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134. Why is electron affinity of Si^{14} is greater than of p^{15} ?

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135. The more electronegative element is :

A. F

B. Cl

C. I

D. Br

Answer: A

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136. Characteristic of transition element is that its ..... Is in complete :

A. d' orbital

B. p' orbital

C. f' orbital

D. s' orbital

Answer: A

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137. The order of decrese in atomic radii for Be, Na, and Mg :

A. Na gt Mg gt Be

B. Mg gt Na gt Be

C. Be gt Na gt Mg

D. Be gt Mg gt Na

Answer: A

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138. Which one has lowest first ionozation potential ?

A. B

B. C

C. N

D. O

Answer: A

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139. Lithium shows diagonal relationship with :

A. Mg

B. Hg

C. Mg and Hg

D. None of these

Answer: A

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140. Which shows highest magnetic moments ?

A. V^{3+}

B. Cr^{3+}

C. Fe^{3+}

D. $C0^{3+}$

Answer: C

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141. Which transition metal ion is diamagnetic ?

A. Co^{2+}

B. Ni^{2+}

C. Cu^{2+}

D. $Zn(2+)$

Answer: D

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142. Which has the smallest size ?

A. Na^+

B. Mg^{2+}

C. Al^{3+}

D. Si^{4+}

Answer: D

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143. The ionization of hydrogen atom would give rise to :

A. Hydride ion

B. Hydronium ion

C. Proton

D. hydroxyl ion

Answer: C

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144. Which ion has the lower magnetic moment ?

A. Cu^{2+}

B. Ni^{2+}

C. Co^{3+}

D. Fe^{2+}

Answer: A

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145. Which elements has the greatest tendancy to lose electrons ?

A. F

B. Fr

C. S

D. Be

Answer: B

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146. Which element has the highest electronegativity ?

A. F

B. H

C. Ne

D. La

Answer: A

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147. The element whose electronic configuration is $1s^2, 2s^2, 2p^6, 3s^2$

A. Metakkoid

B. Metal

C. Nobal gas

D. Non-metal

Answer: B

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148. Which has the smallest size ?

A. Rb

B. K

C. Na

D. Li

Answer: D

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149. Which one is most metallic ?

A. P

B. As

C. Sb

D. Bi

Answer: D

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150. According to modern periodic law , the chemical properties of elements are the periodic functions of there :

A. Density

B. Atomic number

C. Mass number

D. Atomic mass

Answer: B

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151. According to Mendeleev's periodic law the properties of elements were the periodic function of :

A. Atomic weight

B. Atomic number

C. Number of electron

D. None of these

Answer: A

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152. Atomic number of Cr is 24, then Cr^{3+} will be :

A. Diamagnetic

B. Paramagnetic

C. Ferromagnetic

D. None of these

Answer: B

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153. Which set of elements have nearly the same atomic radii ?

A. F, Cl, Br, I

B. Na, K, Rb, Cs

C. Li, Be, B, C

D. Fe, Co, Ni, Cu

Answer: D

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154. Among thr following elements, the most electronegativity is :

A. Oxygen

B. Chlorine

C. Nitrogen

D. Fluorine

Answer: D

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155. Among the following group of element, the one whose element can have positive as well as negative oxidation states are :

A. H, F, O

B. Na, Mg, Al

C. He, Li, Be

D. H, Cl, Br

Answer: D

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156. The process requiring the adsorption of :

A. $F \rightarrow F^-$

B. $H \rightarrow H^-$

C. $Cl \rightarrow Cl^-$

D. $O \rightarrow O^-$

Answer: D

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157. Which of the following is Dobereiner traid

:

A. Cl, Br, I

B. Zn, Cr, Na

C. Ne, Ar, K

D. B, C, Si

Answer: A

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158. In III period, the mst acidic oxide is formed by :

A. Na

B. P

C. Cl

D. S

Answer: C

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159. The example of metalloid elements in the periodic table is :

A. Na and K

B. Cu and Al

C. B and Si

D. Ca and Mg

Answer: C

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160. Manganese belongs to ..... of periodic table:

A. s' block

B. p' block

C. d' block

D. f' block

Answer: C

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161. Which is more electropositive ?

A. Mg

B. Al

C. P

D. S

Answer: A

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162. The elements X,Y,Z and J have the indicated electron configuration starting with the innermost shell. The most metallic element is :

A. X = 2, 8, 3

B. Y = 2, 8, 8

C. Z = 2, 8, 8, 1

D. J = 2, 8, 8, 7

Answer: C

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163. The correct order of the size of the iodine species is :

A. I gt I^+ gtI^-

B. $I > I^- > I^+$

C. $I^+ > I^- > I$

D. $I^- > I > I^+$

Answer: D

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164. An example of a non-stoichiometric compound is :

A. Al_2O_3

B. Fe_3O_4

C. NiO

D. PbO

Answer: B

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165. Which of the following is paramagnetic ?

A. O_2^-

B. CN^-

C. CO

D. NO^+

Answer: A

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166. Which arrangement for the three halogens Cl, Br and I is correct in the order of their increasing electron affinity :

A. Cl, Br, I

B. I, Br, Cl

C. Br, Cl, I

D. I, Cl, Br

Answer: B

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167. Almost the following elements the configuration having the highest ionization energy is :

A. [Ne] $3s^2 3p^1$

B. [Ne] $3s^2 3p^3$

C. [Ne] $3s^2 3p^2$

D. [Ar] $3d^{10} 4s^2 4p^3$

Answer: B

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168. Which series of elements would have nearly the same atomic radii ?

A. Fe, Co, Ni

B. Na, K, Rb

C. F, Cl, Br

D. Li, Be, B

Answer: A

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169. Which metal exahibits more than one oxidation state:

A. Na

B. Mg

C. Fe

D. Al

Answer: C

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170. The element with the highest first ionozation potential is :

A. Boron

B. Carbon

C. Nitrogen

D. Oxygen

Answer: C

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171. The Halogen which has the highest electron affinity :

A. Fluorine

B. Chlorine

C. Bromine

D. Iodine

Answer: B

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172. The correct order of second ionization potential of C, N, O and F is :

A. Cgt N gt Ogt F

B. O gt Ngt F gt C

C. O gt F gt N gt C

D. Fgt O gt N gt C

Answer: C

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173. Which one is metalloid ?

A. Tin

B. Germanium

C. Silicon

D. Carbon

Answer: B

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174. Which of the following is not a metal ?

A. Gold

B. Mercury

C. Scandium

D. Selenium

Answer: D

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175. Which transition metal cation has maximum unpaired electrons ?

A. Mn^{2+}

B. Fe^{2+}

C. Co^{2+}

D. Ni^{2+}

Answer: A

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176. Which one of the following is the smallest In size ?

A. N^{3+}

B. O^{2+}

C. F^-

D. Na^+

Answer: D

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177. Which ion is colourless ?

A. Cu^+

B. Ni^{2+}

C. Co^{2+}

D. Si^{2+}

Answer: A

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178. The least stable cation is :

A. Li^{+}

B. K^{+}

C. Al^{2+}

D. Si^{2+}

Answer: C

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179. The polarity is maximum in :

A. N-F

B. C-F

C. O-F

D. F-F

Answer: B

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180. Which of the following has the largest size ?

A. Br

B. Fr

C. I

D. I^-

Answer: B

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181. Which of the following halides is not oxidised by Mno_2 ?

A. F^-

B. Cl^-

C. Br^-

D. I^-

Answer: A

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182. Lithium shows diagonal relationship with :

A. Al

B. P

C. Si

D. S

Answer: D

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183. The diagonal relationship is not shown by

:

A. Li and Mg

B. Be and Al

C. B and Si

D. N and S

Answer: D

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184. Which species has the maximum ionic radius :

A. Na^+

B. O^{2-}

C. F^-

D. Mg^{2+}

Answer: B

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185. In which pair, second atom is larger than the first :

A. Br, Cl

B. Na,Mg

C. Sr, Ca

D. N, P

Answer: D

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186. An element X occurs in short period having configuration ns^2np^1 . The formula and nature of its oxide is :

A. XO_3 ,basic

B. XO_3 , acidic

C. X_2O_3 , amphoteric

D. X_2O_3 , basic

Answer: C

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187. The smallest atom in the following is :

A. Fluorine

B. Chlorine

C. Iodine

D. Bromine

Answer: A

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188. Electron affinity of noble gases is :

A. Very high

B. Low

C. Almost zero

D. High

Answer: C

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189. A neutral atom will have the lowest ionization potential when electronic configuration is :

A. $1s^1$

B. $1s^2 2s^2 2p^6$

C. $1s^2 2s^2 2p^6 3s^1$

D. $1s^2 2s^2 2p^2$

Answer: C

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190. The lightest metal in the periodic table is :

A. H

B. Mg

C. Ca

D. Li

Answer: D

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191. The group in which all the elemets do not have same number of valency electron is :

A. Zero

B. Second

C. First

D. Seventh

Answer: A

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192. The first largest ionisation potential stands for :

A. Li and Mg

B. K

C. Na

D. Rb

Answer: A

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193. O_2^{2-} is isoelectronic with

A. H_2

B. N_2

C. F_2

D. S

Answer: C

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194. Which one of these is basic ?

A. CO_2

B. SnO_2

C. NO_2

D. SO_2

Answer: B

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195. Among the following outermost configurations of transition which shows the highest oxidation state:

A. $3d^3 4s^2$

B. $3d^5 4s^1$

C. $3d^5 4s^2$

D. $3d^6 4s^2$

Answer: C

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196. Electronegtivity of beryllium is approximately equal to that of :

A. Aluminium

B. Boron

C. Magnesium

D. Sodium

Answer: A

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197. Which element exhibits variable oxidation states ?

A. Al

B. Si

C. Fe

D. Zn

Answer: C

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198. The screening effect of 'd' electrone is :

A. Much less than s-electrons

B. Much more than s-electron

C. Equal to s- electrons

D. Equal to p electrons

Answer: A

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199. Maximum group valancy of halogens with respect to oxygen is :

A. One

B. Five

C. Six

D. Seven

Answer: D

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200. The law of octaves is applicable too :

A. B, N, C

B. Be, Mg, Ca

C. Ar, K, Ca

D. Se, Te, As

Answer: C

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201. Which is most reactive chemically ?

A. Mg

B. Ca

C. Sr,

D. Ba

Answer: D

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202. The first ionisation energy is smallest for the atom with electronic configuration :

A. ps^2np^3

B. ns^2np^4

C. ns^2np^5

D. ns^2np^6

Answer: B

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203. Elements of IA group give flame colour due to :

A. Low IP

B. Low M.P

C. Softness

D. One electron in outermost shell

Answer: A

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204. The number of elements in each long period is :

A. 18

B. 10

C. 8

D. 32

Answer: A

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