



CHEMISTRY

BOOKS - OMEGA PUBLICATION

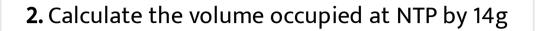
SAMPLE QUESTION PAPER



1. Calculate the volume occupied at NTP by 2.5

mole of CO2





of nitrogen gas.



3. Calculate the mass of 2.5g atom of oxygen



4. Calculate the weight of CO having the same number of oxygen atoms as are present in 22 g of CO2 .

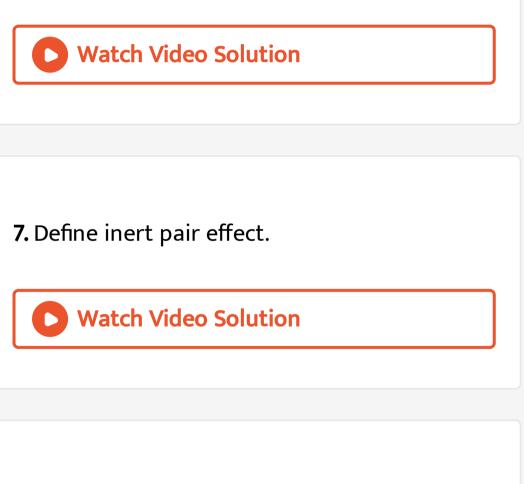


5. Calculate number of molecules in 1ml of CO2

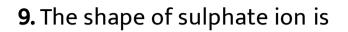
at NTP.

6. Calculate the number of atoms in 3 mol of

NH3.



8. Write the IUPAC name of salicylaldehyde.





10. What are antipyretics ?

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11. Write two main differences between Schottky and Frenkel defect.

12. A solution containing 0.45 g of a urea in 22.5 g of water gave a boiling point elevation of 0.17 K. Calculate the molal elevation constant of water. Molar mass of urea is 60 g mol^{-1} .

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13. Define molality.





14. The units of cell constant are:

 $ohm^{-1}cm^{-1}$, cm, $ohm^{-1}cm$, cm $^{-1}$

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15. Can you store copper sulphate solutions in

a zinc pot?

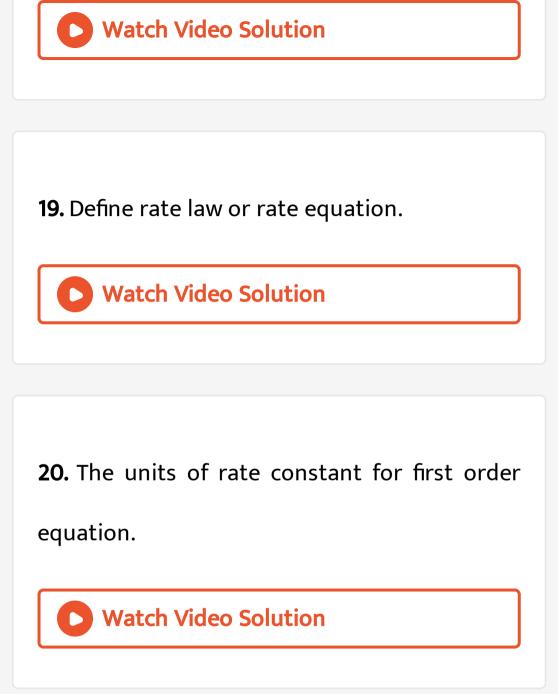
16. State and explain Kohlrausch's law. How would you determine the molar conductance of a weak electrolyte at infinite dilution?

17. What is rusting? What is the chemical formula of rust?

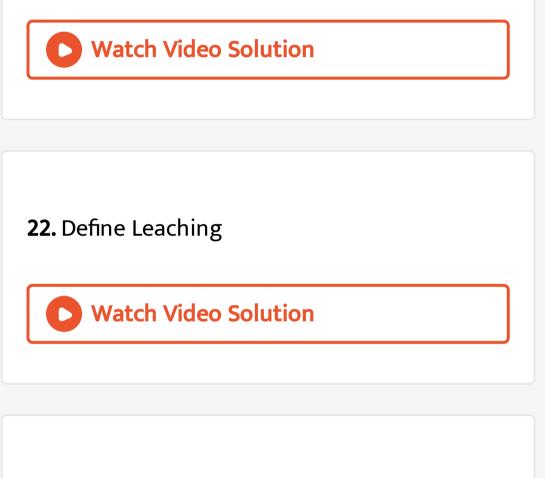
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18. Give one example of zero order reaction.



21. Arrhenius equation is



23. What are micelles ? Give one example of a

micellar system.



24. Explain Brownian movement.

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25. Give the synthesis of following polymers.

Nylon -66



26. Give equation for the following:

 $XeF_6 + H_2O$

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27. What are interhalogen compounds ? Give example.



28. Arrange *HCIO*, *HBrO* and *HIO* in order to decreasing acidic strength giving reasons.



29. Why is water liquid while H_2S a gas ?



30. Write IUPAC name of $[Co(NH_3)_6]Cl_3$.



31. On he basis of valencebond theory explain the structure and magnetic nature of $[Ni(CN)_4]^{2-}$ complex ion.

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32. Explain the following reaction reaction :

Friedal Craft acylation.

33. Write down following name reaction :

Hunsdiecker reaction

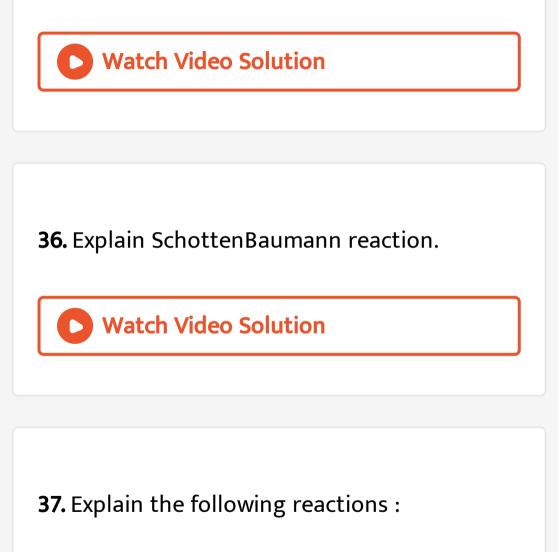
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34. Write the following reactions:

Finkelstein reaction



35. Write esterification reaction.



Kolbe's reaction.

38. Complete the following :

2HCHO+NaOH(50~%~)
ightarrow



39. Complete the following :

 $CH_3COCl + H_2 \xrightarrow{Pd - BaSO_4} \operatorname{S \ or \ quinoline}$



40. Complete the following :

Formaldehyde to urotropine

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41. Complete the following :

Formaldehyde to acetaldehyde

42. Complete the following :

Benzoyl chloride to benzaldehyde.



43. What is Hofmann bromamide degradation

reaction ?



44. Out of ethylamine and ammonia which is

more basic and why?

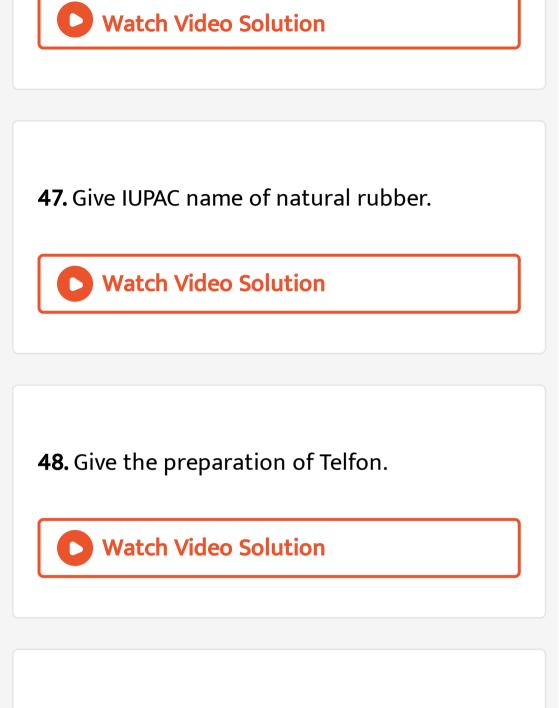
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45. What is mutarotation ?

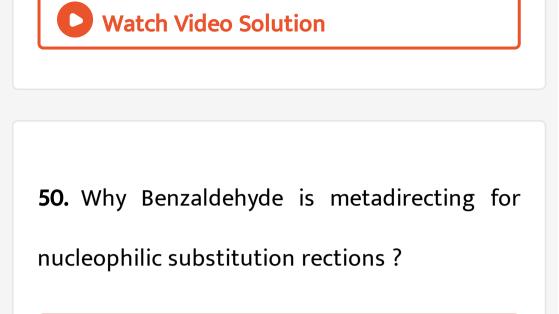
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46. What do you mean by inversion of cane

sugar?

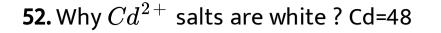


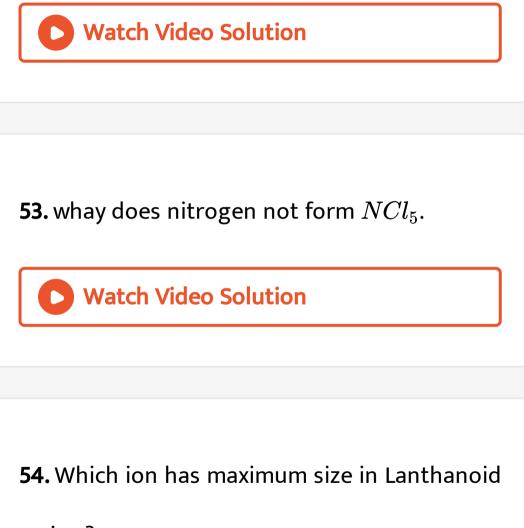
49. Give the name of first antibiotic discovered.



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51. Discuss the shape of P_4O_6 .





series ?





56. What is formalin solution ? Give its one

use..



57. What is Chirality?



58. What happens when phenol is treated with

conc . HNO_3 ?

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59. Close packing is maximum in the crystal which is

A. bcc

B. fcc

C. simple cubic

D. end centered cubic.

Answer:



60. The boiling point of a solvent containing

non volatile solute :

A. is depressed

B. is elevated

C. does not change

D. None of the above

Answer:



61. For a redox reaction to proceed in a cell,

the e.m.f. must be:

A. Positive

B. Negative

C. Fixed

D. zero

Answer:



62. Movement of dispersion medium under the

influence of electric field is known as

A. Electrodialysis

B. Electrophoresis

C. Electroosmosis

D. Cataphoresis

Answer:



63. Zone refining is used for the

A. Concentration of an ore

B. Reduction of metal oxide

C. Purification of metal

D. Purification of an ore.

Answer:

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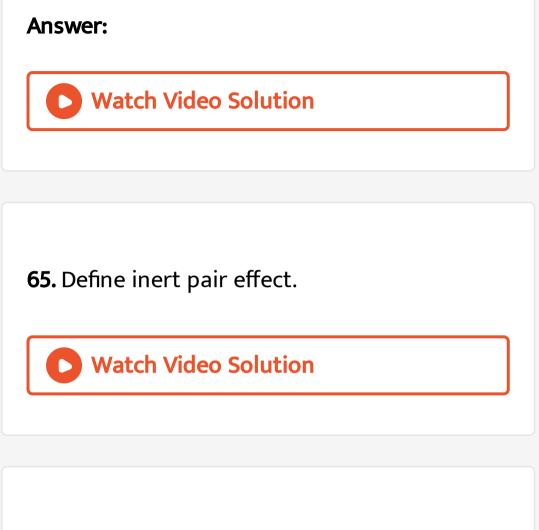
64. Cinnabar is an ore of

A. Hg

B. Zn

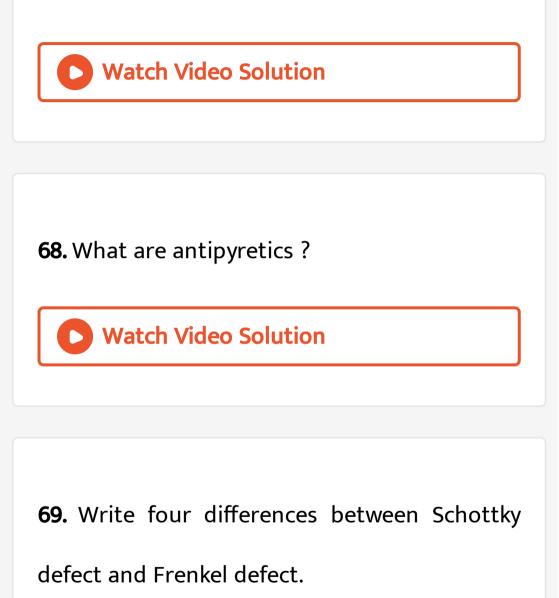
C. Ag

D. Cu



66. Write IUPAC name of isobutyl alcohol.

67. What is the shape of chromate ion ?

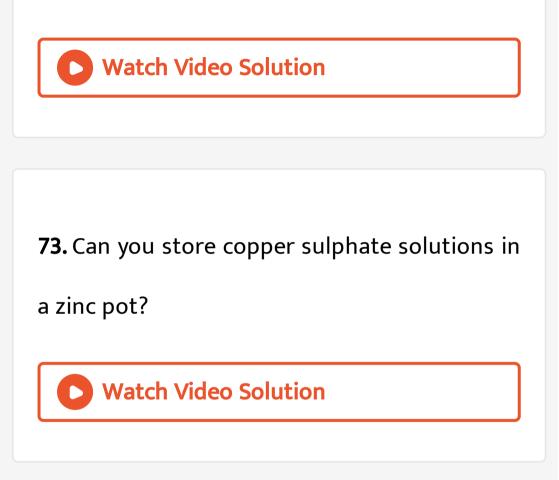


70. A solution containing 12.5 g of a nonelectrolyte substance in 175 g of water gave boiling point elevation of 0.70 K. Calculate the molar mass of the substance (K_b for water = 0.52 K kg mol⁻).

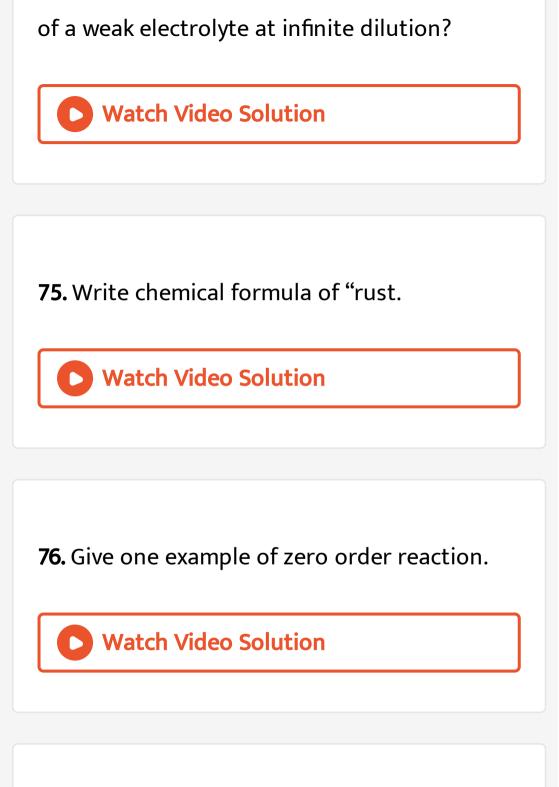
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71. Define molality.

72. Write the units of cell constant.



74. State and explain Kohlrausch's law. How would you determine the molar conductance



77. Define rate law or rate equation.

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78. The units of rate constant for first order equation.

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79. Arrhenius equation is



81. What are micelles ? Give one example of a

micellar system.

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82. Explain Brownian movement.

83. Give the synthesis of following polymers.

Nylon -66



84. Complete the following equation:

 $XeF_6 + H_2O \rightarrow ? + ?$



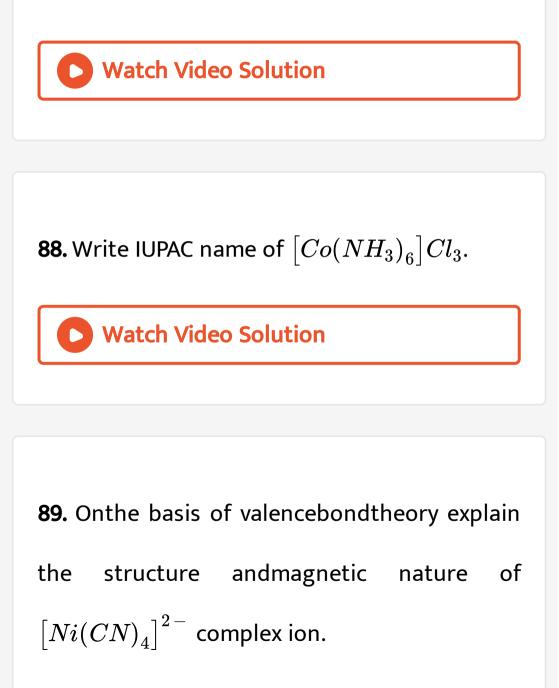
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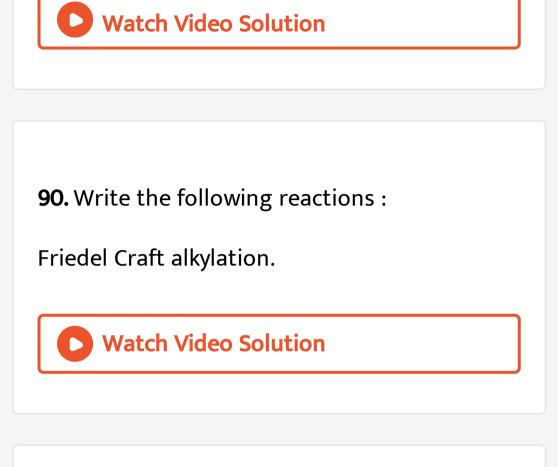
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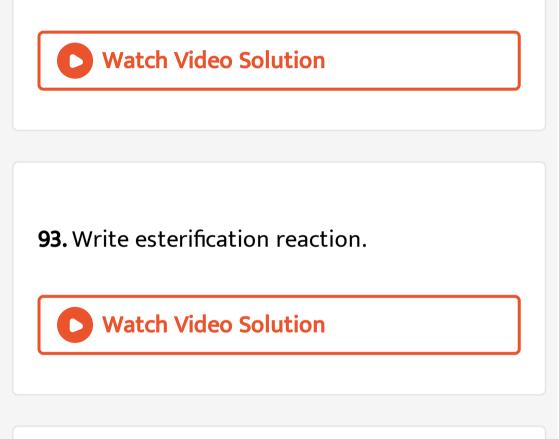


91. Write down following name reaction :

Hunsdiecker reaction

92. Write the following reactions:

Finkelstein reaction



94. Explain SchottenBaumann reaction.

95. Explain the following reactions :

Kolbe's reaction.

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2HCHO+NaOH(50~%~)
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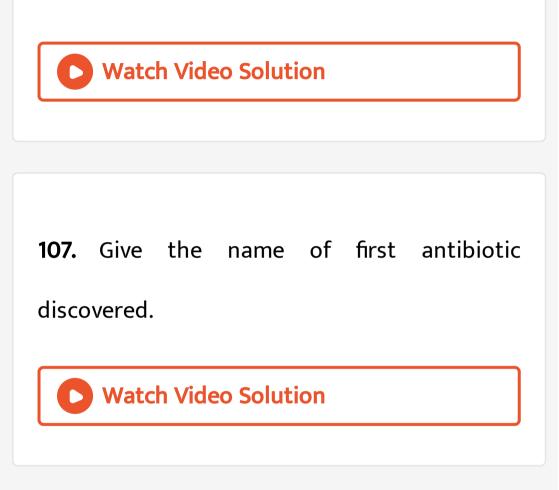
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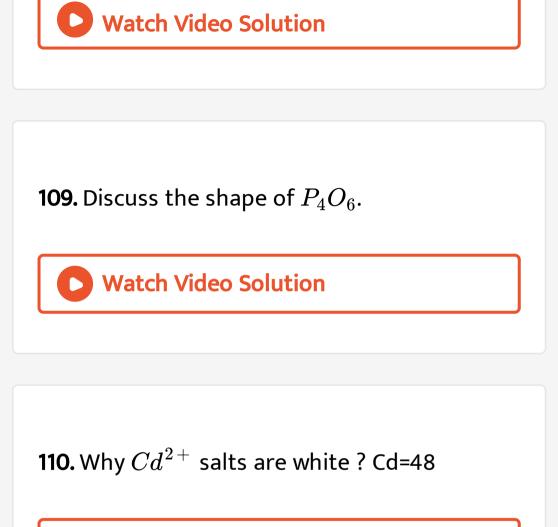
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105. Give IUPAC name of natural rubber.

106. Give the preparation of Telfon.



108. Why Benzaldehyde is metadirecting for nucleophilic substitution rections ?





111. whay does nitrogen not form NCl_5 .

Γ



112. Which ion has maximum size in Lanthanoid series ?

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113. Explain the structure of $Cr_2O_7^{2-}$

114. What is formalin solution ? Give its one

use..

